

## 9-2 Review and Reinforcement

### Chemical Equations

Refer to the unbalanced chemical equation below to answer each of the following questions. Write your answers in the space provided.



1. Write the names of the reactants in this equation.

Aluminum nitrate + Iron (II) chloride

2. Write the names of the products in this equation.

Iron (II) nitrate + aluminum chloride

3. What would the coefficient of  $\text{Al(NO}_3\text{)}_3$  be in the balanced equation? Two

4. What would the coefficient of  $\text{Fe(NO}_3\text{)}_2$  be in the balanced equation? Three

5. What would the coefficient of  $\text{AlCl}_3$  be in the balanced equation? Two

6. How many oxygen atoms would there be on each side of the balanced equation? 18

Answer each of the following questions in the space provided.

7. What should you do if an equation you have written cannot be balanced?

Make sure reactants + products are written correctly, try again, ask for help

8. How is balancing a chemical reaction related to the law of conservation of matter?

mass of reactants must = mass of products

9. Describe the correct way to balance an equation.

add only coefficients to keep charge-balance  
or subscripts

10. How does a chemical equation relate the macroscopic, observable reaction to the interactions occurring at the microscopic level?

Subatomic

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**M Mark**

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