

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the K_a and K_b of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. K_a | 9. K_b |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constants indicated that are associated in pH-related calculations with aqueous acids and bases. Be specific in writing.

1. What substance(s) is/are _____ in water? Hydrogen ion(s) is/are _____.
2. What substance(s) is/are _____ in water? Hydroxide ion(s) is/are _____.
3. What substance(s) is/are _____ in water? Water is/are _____.
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