

19**ACIDS, BASES, AND SALTS****SECTION 19.1 ACID-BASE THEORIES (pages 587–593)**

This section compares and contrasts acids and bases as defined by the theories of Arrhenius, Brønsted-Lowry, and Lewis. It also identifies conjugate acid-base pairs in acid-base reactions.

► Properties of Acids and Bases (pages 587–588)

1. Circle the letters of all the terms that complete the sentence correctly.
The properties of acids include _____.
 - a. reacting with metals to produce oxygen
 - b. giving foods a sour taste
 - c. forming solutions that conduct electricity
 - d. causing indicators to change color
2. Bases are compounds that react with acids to form _____ and a(n) _____.
3. Circle the letters of all the terms that complete the sentence correctly.
The properties of bases include _____.
 - a. tasting bitter
 - b. feeling slippery
 - c. changing the color of an indicator
 - d. always acting as a strong electrolyte

► Arrhenius Acids and Bases (pages 588–590)

4. Match the number of ionizable hydrogens with the type of acid.

_____ one	a. diprotic
_____ two	b. triprotic
_____ three	c. monoprotic
5. Is the following sentence true or false? Only the hydrogens in weak polar bonds are ionizable. _____
6. Hydrogen is joined to a very _____ element in a very polar bond.
7. Alkali metals react with water to produce _____ solutions.

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Disha Experts

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Physical Chemistry for Engineering and Applied Sciences Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto. Designed to be as rigorous as compatible with a first year student's ability to understand the text presents detailed step by step

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