

Chemistry

Unit 2

Practice Exam 1 Answers

Section A—Multiple-choice questions

Question 1

B The high electronegativity difference between hydrogen and oxygen means that the O—H bond in water is highly polar, enabling hydrogen bonds to be formed between nearby water molecules.

Sulfur is less electronegative than oxygen. The polarity of the H—S bond in H_2S is much less than the polarity of the O—H bond in water. The attraction between H_2S molecules is described as a dipole–dipole interaction.

The bond between hydrogen and Te is non-polar. The only interactions between molecules of H_2Te are dispersion forces brought about through the formation of instantaneous dipoles.

Question 2

D The formation of a solution involves breaking bonds between solute particles, breaking intermolecular bonds between solvent molecules, and the formation of bonds between solvent molecules and solute particles.

Question 3

C $n(\text{Na}_2\text{CO}_3) = cV = 0.150 \times \frac{300}{1000} = 0.0450 \text{ mol}$
 $m(\text{Na}_2\text{CO}_3) = mM = 0.0450 \times 106.0 = 4.77 \text{ g}$

Question 4

C Water is a poor conductor of electricity due to the small number of hydronium ions and hydroxide ions that are produced by the self-ionisation reaction of water.

Question 5

A Nitric acid, HNO_3 , is a strong acid because it is completely ionised in water. Ethanoic acid, CH_3COOH , is a weak acid as it is partly ionised in water. A 0.1 M solution is considered to be dilute and 6 M concentrated.

Question 6

B When the acid H_2PO_4^- donates a proton, it forms the conjugate base HPO_4^{2-} .

Question 7

A pH is defined as $\text{pH} = -\log_{10}[\text{H}_3\text{O}^+]$. Since HCl is a strong acid and completely ionised in solution, it can be assumed that in each beaker $[\text{HCl}] = [\text{H}_3\text{O}^+]$. pH is only dependent on the $[\text{H}_3\text{O}^+]$, not the volume of solution. Because the definition involves the negative logarithm, the pH increases as the $[\text{H}_3\text{O}^+]$ decreases. The solution that has the lowest $[\text{H}_3\text{O}^+]$ will have the highest pH.

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Yuli Rahmawati, Peter Taylor

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Physical Chemistry Calculations Rodney J. Sime,2005 Physical Chemistry Calculations is a practical guide for students and instructors who want to learn how to use the most popular spreadsheet and computational software to solve problems in physical chemistry The book provides students with a complementary approach to the chemistry and physics they are learning in the classroom Physical Chemistry Calculations also gives a solid introduction to calculations with Excel VB VBA MathCad and Mathematica Research in Chemistry Education Liliana Mammino,Jan Apotheker,2021-05-17 This volume emphasizes the role of chemical education for development and in particular for sustainable development in Africa by sharing experiences among specialists across the African continent and with specialists from other continents It considers all areas and levels of chemistry education gives specific attention to known major challenges and encourages explorations of novel approaches The chapters in this book describe new teaching approaches approach explorations and in class activities analyse educational challenges and possible ways of addressing them and explore cross discipline possibilities and their potential benefits for chemistry education This makes the volume an up to date compendium for chemistry educators and educational researchers worldwide Miscellaneous Topic for Higher Mathematics: Course in Mathematics for the IIT-JEE and Other Engineering Entrance Examinations K.R. Choubey, Ravikant Choubey, Chandrakant Choubey, Whitaker's Five-year Cumulative Book List ,1968

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