

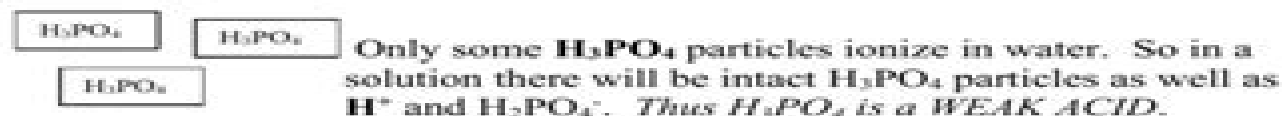
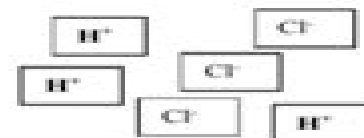
Acid and Base Strength Worksheet

Name _____

Period _____ Date _____

What makes an acid *acidic* are the *excess hydrogen ions (H^+)** that these substances cause in a water solution.

Strong acids, like HCl, dissociate (break up) completely in solution, giving a maximum number of hydrogen ions. Another way to say this is, strong acids **IONIZE** completely while weak acids don't.



So, how do you know which acids are strong and which are weak?
Easy; there are a few rules.

First, a few definitions:

- A **binary acid** is made of **two elements**: hydrogen and something else. HCl is a binary acid made of hydrogen and chloride. Count 'em. Hydrogen (one), chloride (two). Ta-da!
- A **ternary acid** is made of **more than two elements**: hydrogen and at least two other elements. H_2SO_4 , with three elements, is a ternary acid. So is H_3PO_4 and acetic acid, $HC_2H_3O_2$. (Notice acid formulas always start with a hydrogen. These are the hydrogens that can break free in water to a lesser or greater degree depending on the strength of the acid.)

THE RULES for deciding if an acid is strong or weak!

- **The only strong binary acids are HCl, HBr, and HI.** (Notice Cl, Br, and I are together on the periodic table.) All other binary acids are weak.
- Ternary acids are a little trickier, but not much. Compare the number of oxygens and hydrogens in the formula. **If the number of oxygens exceeds the number of hydrogens by 2 or more, the acid is strong.** All other ternary acids are weak.

H_3PO_4 3 hydrogens and 4 oxygens (only one more oxygen than hydrogen). **WEAK**

H_2SO_4 2 hydrogens and 4 oxygens (two more oxygens than hydrogen). **STRONG.**

*Technically, the free H^+ usually combines with a water molecule to form a hydronium ion (H_3O^+), but we will ignore this for the time being.

19 3 Strengths Of Acids And Bases Worksheet Answers

Brian G. Cox



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Table of Contents 19 3 Strengths Of Acids And Bases Worksheet Answers

1. Understanding the eBook 19 3 Strengths Of Acids And Bases Worksheet Answers
 - The Rise of Digital Reading 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Strengths Of Acids And Bases Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Personalized Recommendations
 - 19 3 Strengths Of Acids And Bases Worksheet Answers User Reviews and Ratings
 - 19 3 Strengths Of Acids And Bases Worksheet Answers and Bestseller Lists

5. Accessing 19 3 Strengths Of Acids And Bases Worksheet Answers Free and Paid eBooks
 - 19 3 Strengths Of Acids And Bases Worksheet Answers Public Domain eBooks
 - 19 3 Strengths Of Acids And Bases Worksheet Answers eBook Subscription Services
 - 19 3 Strengths Of Acids And Bases Worksheet Answers Budget-Friendly Options
6. Navigating 19 3 Strengths Of Acids And Bases Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 3 Strengths Of Acids And Bases Worksheet Answers Compatibility with Devices
 - 19 3 Strengths Of Acids And Bases Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Highlighting and Note-Taking 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Interactive Elements 19 3 Strengths Of Acids And Bases Worksheet Answers
8. Staying Engaged with 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Strengths Of Acids And Bases Worksheet Answers
9. Balancing eBooks and Physical Books 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Strengths Of Acids And Bases Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Setting Reading Goals 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Fact-Checking eBook Content of 19 3 Strengths Of Acids And Bases Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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