

Name: \_\_\_\_\_ per: \_\_\_\_\_

### Worksheet - Reaction Rates

Use this reaction for the questions below:  $C_6H_{12}O_6 (s) + 6 O_2(g) \rightarrow 6 H_2O (g) + 6 CO_2 (g)$

- What happens to the concentrations of:
  - $C_6H_{12}O_6$  &  $O_2$  as the reaction proceeds  $\rightarrow$ ?
  - $H_2O$  +  $CO_2$  as the reaction proceeds  $\rightarrow$ ?
- According to the collision theory, what 3 circumstances are needed for  $C_6H_{12}O_6$  &  $O_2$  to react?
- What is the activation energy for a chemical reaction?
- Use the equation & the collision theory to explain:



Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a. <u>Increasing the temperature</u>	Ex: Increase (springs up)	Ex: Molecules move faster & collide more = increased rxn rate
b. <u>Increasing the concentration</u> of $C_6H_{12}O_6$		
c. <u>Decreasing the concentration</u> of $O_2$		
d. <u>Increase the surface area</u> by chewing up food in your mouth.		
e. <u>Decreasing the temperature</u>		
f. <u>Increasing the pressure</u> in the container		
g. <u>Decreasing the concentration</u> of $H_2O$		
h. <u>Increasing the volume</u> of the container the reaction occurs in		
i. <u>Increasing the concentration</u> of $CO_2$		
j. <u>Using a catalyst</u> (like salivary amylase)		

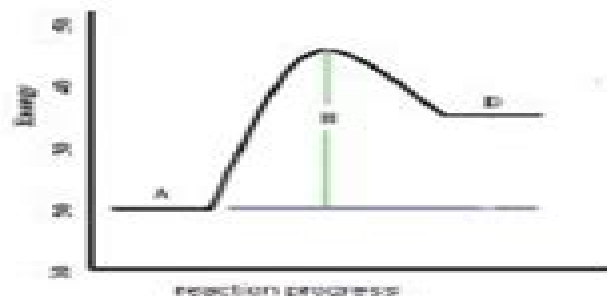
- On the accompanying energy diagram, label the following terms:

a. reactants    b. products    c. activation energy

- On the graph to the right, draw and label what this diagram would look like if a catalyst was added to the reaction.

- Graph reading

- How much energy (kJ) do the reactants have?
- How much energy (kJ) do the products have?
- How much energy (kJ) is required to activate this un-catalyzed reaction?
- Is this reaction endothermic or exothermic? How do you know?



# 18 Reaction Rates And Equilibrium Worksheet Answers

**L Darling-Hammond**



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Software for Teaching Science Roger Frost,1998      *Chemical Kinetics with Mathcad and Maple* Viktor Korobov,Valery Ochkov,2011-05-26 The authors explain at length the principles of chemical kinetics and approaches to computerized calculations in modern software suites mathcad and maple Mathematics is crucial in determining correlations in chemical processes and requires various numerical approaches Often significant issues with mathematical formalizations of chemical problems arise and many kinetic problems can't be solved without computers Numerous problems encountered in solving kinetics calculations with detailed descriptions of the numerical tools are given Special attention is given to electrochemical reactions which fills a gap in existing texts not covering this topic in detail The material demonstrates how these suites provide quick and precise behavior predictions for a system over time for postulated mechanisms Examples i e oscillating and non isothermal reactions help explain the use of mathcad more efficiently Also included are the results of authors own research toward effective computations      *Books in Print Supplement* ,2002      **Radioactive Waste Management** ,1981

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*Rates of Reaction and Equilibria* Clifford Othen,1972      *Chemical Equilibria in Solution* M. J. Blandamer,1992      **A**

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