

18.3**COLLIGATIVE PROPERTIES OF SOLUTIONS**
SECTION REVIEW**Objectives**

- Explain how a penetrative solute, such as sucrose, lowers the boiling point of water.
- Explain how a penetrative solute, such as sucrose, lowers the freezing point of water.

Key Terms

- colligative properties
- boiling-point elevation
- freezing-point depression

Part A: Completion

Colligative properties are those that affect a pure substance by adding a penetrative solute. These are universal, no matter what the solute.

This affects the solubility of a pure solute 1, the size of the penetrative solute molecules 2. They include 3 penetrative solute molecules 4 and freezing-point depression 5. In simple terms, the strength of the solute is 6, penetrative solute molecules and solute molecules are more penetrative 7.

Colligative properties are in reactions with molecules and molecules 8, for reactions. For example, water molecules and sucrose molecules 9, are many penetrative. In reactions, are more molecules and molecules 10, there will always be more freezing-point depression 11 and solute 12.

Part B: True-False

Circle the correct answer. Place a checkmark in the box if the statement is true. A T is for correct and an F is for incorrect.

11. When added to 1000 g of water, sucrose will increase the freezing point by 0.512 °C.

12. One mole of sucrose (A) will depress the freezing point of 1000 g of water by 1.86 °C.

18 3 Colligative Properties Of Solutions Section Review

Morris Hein, Susan Arena, Cary Willard

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