

18.3**COLLIGATIVE PROPERTIES OF SOLUTIONS****SECTION REVIEW****Objectives**

- Explain on a particle basis why a solution has a lower vapor pressure than the pure solvent of that solution.
- Explain on a particle basis why a solution has an elevated boiling point and a depressed freezing point compared with the pure solvent.

Key Terms

- colligative properties
- boiling-point elevation
- freezing-point depression

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or even line.

The effects in solution of a nonvolatile _____ on the properties of the solvent are called _____. They include _____ point and vapor pressure _____ and boiling point _____. In each case, the magnitude of the effect is _____ proportional to the number of solute molecules or ions present in the _____.

Colligative properties are a function of the number of solute _____ in solution. For example, one mole of sodium chloride produces _____ as many particles in solution as one mole of sucrose and, thus, will depress the freezing point of water _____ as much.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

Part B True-False

Classify each of these statements as an always true, AT; sometimes true, ST; or never true, NT.

- _____ 11. When added to 1000 g of water, 2 moles of a solute will decrease the boiling point by 0.212°C .
- _____ 12. One mole of solute A will depress the freezing point of 1000 g of water the same as one mole of solute B.

18 3 Colligative Properties Of Solutions Section Review

Morris Hein,Susan Arena,Cary Willard



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Table of Contents 18 3 Colligative Properties Of Solutions Section Review

1. Understanding the eBook 18 3 Colligative Properties Of Solutions Section Review
 - The Rise of Digital Reading 18 3 Colligative Properties Of Solutions Section Review
 - Advantages of eBooks Over Traditional Books
2. Identifying 18 3 Colligative Properties Of Solutions Section Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 18 3 Colligative Properties Of Solutions Section Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from 18 3 Colligative Properties Of Solutions Section Review
 - Personalized Recommendations
 - 18 3 Colligative Properties Of Solutions Section Review User Reviews and Ratings
 - 18 3 Colligative Properties Of Solutions Section Review and Bestseller Lists
5. Accessing 18 3 Colligative Properties Of Solutions Section Review Free and Paid eBooks

- 18 3 Colligative Properties Of Solutions Section Review Public Domain eBooks
- 18 3 Colligative Properties Of Solutions Section Review eBook Subscription Services
- 18 3 Colligative Properties Of Solutions Section Review Budget-Friendly Options
- 6. Navigating 18 3 Colligative Properties Of Solutions Section Review eBook Formats
 - ePub, PDF, MOBI, and More
 - 18 3 Colligative Properties Of Solutions Section Review Compatibility with Devices
 - 18 3 Colligative Properties Of Solutions Section Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 18 3 Colligative Properties Of Solutions Section Review
 - Highlighting and Note-Taking 18 3 Colligative Properties Of Solutions Section Review
 - Interactive Elements 18 3 Colligative Properties Of Solutions Section Review
- 8. Staying Engaged with 18 3 Colligative Properties Of Solutions Section Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 18 3 Colligative Properties Of Solutions Section Review
- 9. Balancing eBooks and Physical Books 18 3 Colligative Properties Of Solutions Section Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 18 3 Colligative Properties Of Solutions Section Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 18 3 Colligative Properties Of Solutions Section Review
 - Setting Reading Goals 18 3 Colligative Properties Of Solutions Section Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 18 3 Colligative Properties Of Solutions Section Review
 - Fact-Checking eBook Content of 18 3 Colligative Properties Of Solutions Section Review
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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










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