

## Section Review

## Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium.
- Identify three stresses that can change the equilibrium position of a chemical system.
- Explain what the value of  $K_{\text{eq}}$  indicates about the position of equilibrium.

## Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Chatelier's principle
- equilibrium constant ( $K_{\text{eq}}$ )

## Key Equation

$$K_{\text{eq}} = \frac{[\text{C}]^c \times [\text{D}]^d}{[\text{A}]^a \times [\text{B}]^b}$$

When  $a\text{A} + b\text{B} \rightleftharpoons c\text{C} + d\text{D}$

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short sentence, or number.

## 18 3 Reversible Reactions And Equilibrium Answer Key

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Singhal Atul, 2011-09 **How to Pass Dental Aptitude Tests, Questions and Answers** College Publishing Corporation, Brooklyn, 1962

**Cell and Molecular Biology** Gerald Karp, 2009-10-19 Karp continues to help biologists make important connections between key concepts and experimentation The sixth edition explores core concepts in considerable depth and presents experimental detail when it helps to explain and reinforce the concepts The majority of discussions have been modified to reflect the latest changes in the field The book also builds on its strong illustration program by opening each chapter with VIP art that serves as a visual summary for the chapter Over 60 new micrographs and computer derived images have been added to enhance the material Biologists benefit from these changes as they build their skills in making the connection

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