

Section Review

Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium
- Identify three stresses that can change the equilibrium position of a chemical system
- Explain what the value of K_{eq} indicates about the position of equilibrium

Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Châtelier's principle
- equilibrium constant (K_{eq})

Key Equation

$$K_{eq} = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

When $aA + bB \rightleftharpoons cC + dD$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

1. A reversible reaction is one in which the reaction can proceed in both directions. The reaction is reversible if the reaction is reversible.

2. The equilibrium constant, K_{eq} , is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its coefficient in the balanced chemical equation. The value of K_{eq} is constant for a given reaction at a given temperature.

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18 2 Reversible Reactions And Equilibrium Worksheet Answers

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