

Section Review

Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium
- Identify three stresses that can change the equilibrium position of a chemical system
- Explain what the value of K_{eq} indicates about the position of equilibrium

Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Châtelier's principle
- equilibrium constant (K_{eq})

Key Equation

$$K_{eq} = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

When $aA + bB \rightleftharpoons cC + dD$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

1. A reversible reaction is one in which the reaction can proceed in both directions.

2. A reversible reaction is one in which the reaction can proceed in both directions.

3. A reversible reaction is one in which the reaction can proceed in both directions.

4. A reversible reaction is one in which the reaction can proceed in both directions.

5. A reversible reaction is one in which the reaction can proceed in both directions.

6. A reversible reaction is one in which the reaction can proceed in both directions.

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11. A reversible reaction is one in which the reaction can proceed in both directions.

12. A reversible reaction is one in which the reaction can proceed in both directions.

13. A reversible reaction is one in which the reaction can proceed in both directions.

14. A reversible reaction is one in which the reaction can proceed in both directions.

15. A reversible reaction is one in which the reaction can proceed in both directions.

16. A reversible reaction is one in which the reaction can proceed in both directions.

17. A reversible reaction is one in which the reaction can proceed in both directions.

18. A reversible reaction is one in which the reaction can proceed in both directions.

19. A reversible reaction is one in which the reaction can proceed in both directions.

20. A reversible reaction is one in which the reaction can proceed in both directions.

18 2 Reversible Reactions And Equilibrium Answers

Nicholas R. Romano, Philip K. Cameron



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Modern Chemistry Raymond E. Davis, 1999 2000 2005

State Textbook Adoption Rowan Salisbury

British Chemical and Physiological Abstracts, 1950

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Applied Science & Technology Index, 1974

Russian Journal of Physical Chemistry, 2000

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

Russian Journal of Inorganic Chemistry, 1989

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