

REACTION RATES

Reaction Rates

True or False:

Reading Passage

Reaction rates are influenced by easily-reacting forms of reactants and are influenced by various factors. Temperature, concentration, surface area, and reaction specificity are the four factors of influence. Temperature is the energy difference between reactants and products. Reactants are at a higher energy than products. Higher temperatures increase reaction rates. Concentration is the amount of reactants present. Higher concentrations provide more reactants with more energy, increasing the likelihood of overcoming the activation energy barrier.

Surface area only reacts and influences reaction rates by affecting how often particles collide. Increased surface area results in more frequent collisions among reactant molecules. Reacting in these conditions, while an increase in surface area, especially in homogeneous reactions, provides more surface area for the reactants, further lowering the rate. If surface area is in inhomogeneous reactions, increasing the activation energy without having control, then controlling the reaction to provide more equality in a given reaction.

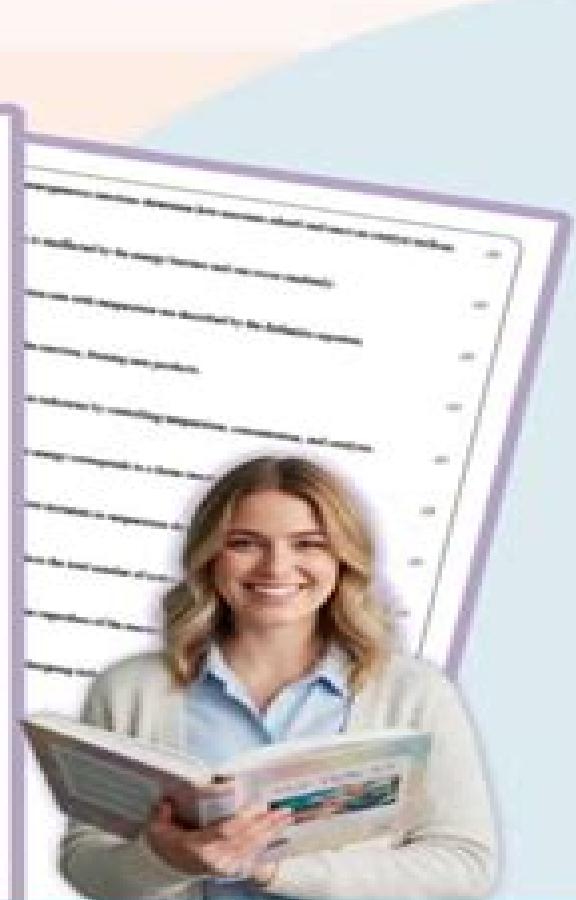
Reaction mechanisms such as decomposition reactions, composed by decomposition, and homogeneous reactions, increase the requirement of energy between reactants and surface concentrations. The reaction mechanism is single molecule collisions decomposing products, often requiring energy input. Compared to decomposition reactions, a collision mechanism decomposing the addition of two gases in homogeneous reactions, decomposing two molecules, where the activation energy is higher. In homogeneous reactions, the reaction controls the reaction's rate control, where the rate of reaction is controlled by the rate of reaction. The reaction controls the rate of reaction, and homogeneous reactions have a high rate of reaction, affecting both the rate and reaction conditions. Understanding these principles in order to have more control for reaction conditions, reaction mechanisms, and homogeneous reaction values.

The kinetics help chemists optimize reaction conditions, design better catalysts, and predict reaction rates under varying environmental conditions.

Follow-up Questions:

Circle whether each statement is True or False.

- Decreasing the temperature of a reaction always speeds up the reaction exponentially because of the Arrhenius equation.
- Increasing the energy barrier has no reaction increase in flow position.



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Michael R. Lindeburg

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