

REACTION RATES AND EQUILIBRIUM

SECTION 18.1 RATES OF REACTION (pages 545–547)

The reaction coordinate below is typical for the first step of a chemical reaction. In which stage is the reaction slowest? Use *Chemical Change* after each step to show how the reaction is affected by the reaction conditions.

► Collision Theory (pages 541–542)

1. How are rates of chemical change expressed?

Chemical change—reaction rate—represents the amount of reactants or products changing over time.

2. Look at Figure 18.1 on page 542. Is a typical reaction, no matter what, the product of **interactions** decreases and the product of **collisions** increases?

Chemical change—reaction rate—represents the amount of reactants changing over time.

3. What does collision theory say about the outcome of reaction, such as mechanism involving the reactants, products, and energy?

Chemical change—reaction rate—represents the amount of reactants changing over time.

4. Look at the figure below. One element is collision slow, resulting in slow reaction rate. Label it **efficiency** problem. Label the other collisions as **inefficiency** problems.



5. In the following sentence, true or false? Particles having the necessary kinetic energy to react have a greater chance of colliding when they collide. **True**

6. Look at Figure 18.3 on page 543. Which arrangement of reactants indicates that **kinetic energy**?

a. reactants

b. products

c. products

18 1 Rates Of Reaction Answer

**Thirumalaisamy P.
Velavancorresponding**

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