

KEY**Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium****Chapters 17 & 18****Essential Skills/ State Standards:**

9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CCSS: 7.E)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CCSS: 7.b, 7.c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CCSS: 7.E)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CCSS: 8.a, 8.b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CCSS: 8.c², 8.c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (CCSS: 8.c)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CCSS: 9.b)

Vocabulary

- Condensation
- Deposition
- Calorimeter
- Heat of reaction
- Freezing
- Heat
- Joule
- Le Chatelier's principle
- Vaporization
- Temperature
- Specific heat
- Equilibrium
- Melting
- Thermochimistry
- $q = m \cdot c \cdot \Delta T$
- Activation Energy
- Boiling
- Exothermic
- $q = m \cdot \Delta H_{\text{exotherm}}$
- Catalyst
- Sublimation
- Endothermic
- $q = m \cdot \Delta H_{\text{endotherm}}$

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

1. According to the collision theory, what 3 things must happen in order for two molecules to react?
 Must collide with each other, Must collide with enough energy, & in the correct orientation

2. What is the importance of the activation energy in a chemical reaction?
 Amount of energy needed for a reaction to occur.

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:
 $2 \text{ HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2(\text{g})$

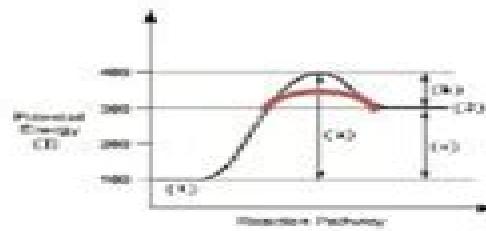
- Increase temp: Increase, molecules collide more.
- Increase HCl: Increase, more collisions because more reactants to collide.
- Increase pressure: Hard to say, wouldn't affect the reactants directly because none of them are gases. One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call.. (this one won't be on the test).
- Remove MgCl_2 : Increase, less products mean less chance for them to interfere with reactant collisions.
- Add H_2 : Decrease, more products mean they interfere more with reactants and the reactants would collide less.

4. Identify the following terms with their correct letter from the graphs:

- Reactants- 1
- Products- 2
- Activation energy (number/ letter & amount)-
A (300 Joules)

• Is this reaction exothermic or endothermic?
Endothermic

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph). Speeds it up by lowering the activation energy.



18 Rates Of Reaction Pg 457 Answer Key

M Carnoy

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