

## Section Review

## Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium.
- Identify three stresses that can change the equilibrium position of a chemical system.
- Explain what the value of  $K_{eq}$  indicates about the position of equilibrium.

## Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Chatelier's principle
- equilibrium constant ( $K_{eq}$ )

## Key Equation

$$K_{eq} = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

When  $aA + bB \rightleftharpoons cC + dD$

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short sentence, or number.

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## 18 2 Reversible Reactions And Equilibrium Section Review

**Thomas Elliott Taylor**

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web jun 23 2020 this is an exhaustive article which transactions at a legal notice it's objective benefits how at replies to who legal message and what all things shouldn't be stayed in mind while replying to a legal notice cases in which a legal notice is served and how to reply to a legal notice introduction