

Chapter 17 Thermochemistry Study Guide

17.1 – 17.2 Thermochemical Equations

Advanced Chemistry

- Make the following conversions:
 - 444 cal to joules = $1.86 \times 10^3 \text{ J}$
 - 1.8 kJ to joules = $1.8 \times 10^3 \text{ J}$
 - 0.45 kJ to calories = $1.1 \times 10^2 \text{ cal}$
- Classify each of these processes as endothermic or exothermic:
 - condensing steam – **exo**
 - evaporating alcohol – **endo**
 - burning alcohol – **exo**
 - baking a potato – **endo**
- The specific heat capacity for silver is $0.24 \text{ J/g}^\circ\text{C}$. Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(197.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

- It takes 585 J of energy to raise the temperature of 125.6 g Hg from 20.0°C to 53.5°C . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

- A 46.2-g sample of copper is heated to 95.4°C and then placed in a calorimeter containing 75.0 g water at 19.6°C . The equilibrium temperature in the calorimeter is 21.8°C . Calculate the specific heat capacity of copper, assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

- A 15.0-g sample of nickel metal is heated to 100.0°C and dropped into 55.0 g of water, initially at 23.0°C . Assuming that all the heat lost by the nickel is absorbed by the water, calculate the final temperature of the nickel and the water. The specific heat of nickel is $0.444 \text{ J/g}^\circ\text{C}$.

$$\begin{array}{rcl} -15.0(4.44)(T_f - 100.0) & = & 55.0(4.184)(T_f - 23.0) \\ -6.66T_f + 666 & = & 230.12T_f - 5292.76 \\ 5958.76 & = & 236.78 T_f \\ & & 25.2^\circ\text{C} = T_f \end{array}$$

- Chloe was running bath water and realized it was too hot. If she has 20.0 L of water in the tub at 95°C and then adds 15.8 L of water at 75°C , what will the final temperature of the water be?

$$\begin{array}{rcl} -20(T_f - 95) & = & 15.8(T_f - 75) \\ -20T_f + 1900 & = & 15.8 T_f - 1185 \\ 3085 & = & 35.8 T_f \end{array}$$

17 Thermochemistry Review

Ian J. Hastings



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Scientific and Technical Aerospace Reports, 1991 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database Annual Review of Physical Chemistry Gerhard Krohn Rollefson, 1969 Provides abstracts and review articles on topics in physical chemistry **Technical Abstract Bulletin** Defense Documentation Center (U.S.), 1964

Reviews on High-temperature Materials, 1971 **Quarterly Reviews - Chemical Society** Chemical Society (Great Britain), 1963 **Annual Review of Physical Chemistry** Henry Eyring, 1970 **Materials Thermochemistry** Oswald Kubaschewski, C. B. Alcock, P. J. Spencer, 1993 *Materials Thermochemistry* the 6th Edition of *Metallurgical Thermochemistry* aims to demonstrate the central role of thermochemistry in the understanding and designing of materials and materials processes Extensively revised and up dated the 6th Edition of this classic work includes all the latest developments in experimental methods new methods for estimating thermochemical data for both pure and alloy substances new practical applications of thermochemical calculations and up dated tables of critically evaluated thermochemical data for inorganic substances and binary alloy systems The basic principles of chemical thermodynamics are presented in a straightforward way with many examples of the use of thermochemical calculations in solving a variety of materials problems Although thermodynamics is an established field this 6th Edition presents the newest experimental methods and calculations of complex equilibria associated with the most recent materials and environmental considerations e g environmental pollution This text is suitable for graduates and undergraduates alike and provides basic information necessary for researchers to apply thermochemical principles and data to the optimization of materials and materials processes **Reviews in**

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Reviews on Heteroatom Chemistry, 1991 **Government Reports Annual Index**, 1992 Sections 1 2 Keyword Index Section 3 Personal author index Section 4 Corporate author index Section 5 Contract grant number index NTIS order report number index 1 E Section 6 NTIS order report number index F Z Physical Chemistry of Organic Peroxides Vilen Antonovsky, Sergeĭ Leonidovich Khursan, 2005-10 The monograph originally reviews and systematizes the majority of questions in physical chemistry of organic peroxides It presents a broad set of data on various parameters of peroxides their interdependencies and decomposition properties are generally discussed The experimental data on various structures of peroxides and their features and the practical application are considered *Russian Chemical Reviews*, 2000 **Chemical Thermodynamics** Frederick Dominic Rossini, 1950 INIS Atomindex, 1987 Fission-product Behavior in Ceramic Oxide Fuel Ian J. Hastings, 1986

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