

# Chapter 17 Thermochemistry Study Guide

## 17.1 – 17.2 Thermochemical Equations

### Answer Key

1. Make the following conversions:

- 444 cal to joules =  $1.86 \times 10^3$  J
- 1.8 kJ to joules =  $1.8 \times 10^3$  J
- 0.45 kJ to calories =  $1.1 \times 10^3$  cal

2. Classify each of these processes as endothermic or exothermic:

- condensing steam = **exo**
- burning alcohol = **exo**
- evaporating alcohol = **endo**
- baking a potato = **endo**

3. The specific heat capacity for silver is  $0.24 \text{ J/g}^\circ\text{C}$ . Calculate the energy required to raise the temperature of 150.0 g Ag from  $273^\circ\text{C}$  to  $298^\circ\text{C}$ . Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(0.24)(298-273) = 9.0 \times 10^3 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(167.87 \text{ g/mol Ag}) = 26.2 \text{ J/mol}$$

4. It takes 565 J of energy to raise the temperature of 125.0 g Hg from  $20.0^\circ\text{C}$  to  $53.5^\circ\text{C}$ . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta T = 565/(125.0)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

5. A 46.2-g sample of copper is heated to  $95.4^\circ\text{C}$  and then placed in a calorimeter containing 75.0 g water at  $19.6^\circ\text{C}$ . The equilibrium temperature in the calorimeter is  $21.8^\circ\text{C}$ . Calculate the specific heat capacity of copper, assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

6. A 15.0-g sample of nickel metal is heated to  $100.0^\circ\text{C}$  and dropped into 55.0 g of water, initially at  $23.0^\circ\text{C}$ . Assuming that all the heat lost by the nickel is absorbed by the water, calculate the final temperature of the nickel and the water. The specific heat of nickel is  $0.444 \text{ J/g}^\circ\text{C}$ .

$$-15.0(0.444)(T_f - 100.0) = 55.0(4.184)(T_f - 23.0)$$

$$-6.66T_f + 666 = 230.12T_f - 5292.76$$

$$5958.76 = 236.76 T_f$$

$$25.2^\circ\text{C} = T_f$$

7. Chloe was running bath water and realized it was too hot. If she has 20.0 L of water in the tub at  $95^\circ\text{C}$  and then adds 15.0 L of water at  $75^\circ\text{C}$ , what will the final temperature of the water be?

$$-20(T_f - 95) = 15.0(T_f - 75)$$

$$-20T_f + 1900 = 15.0 T_f - 1125$$

$$3025 = 35.0 T_f$$

# 17 Thermochemistry Review

**Ian J. Hastings**

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**Reviews on High-temperature Materials**, 1971    **Quarterly Reviews - Chemical Society** Chemical Society (Great Britain), 1963    **Annual Review of Physical Chemistry** Henry Eyring, 1970    **Materials Thermochemistry** Oswald Kubaschewski, C. B. Alcock, P. J. Spencer, 1993 Materials Thermochemistry the 6th Edition of Metallurgical Thermochemistry aims to demonstrate the central role of thermochemistry in the understanding and designing of materials and materials processes Extensively revised and up dated the 6th Edition of this classic work includes all the latest developments in experimental methods new methods for estimating thermochemical data for both pure and alloy substances new practical applications of thermochemical calculations and up dated tables of critically evaluated thermochemical data for inorganic substances and binary alloy systems The basic principles of chemical thermodynamics are presented in a straightforward way with many examples of the use of thermochemical calculations in solving a variety of materials problems Although thermodynamics is an established field this 6th Edition presents the newest experimental methods and calculations of complex equilibria associated with the most recent materials and environmental considerations e.g. environmental pollution This text is suitable for graduates and undergraduates alike and provides basic information necessary for researchers to apply thermochemical principles and data to the optimization of materials and materials processes    **Reviews in**

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