

ANSWER KEY: Ch. 17 Acid-base and Solubility Equilibria

Section 17.1

1. **Acid** is a substance that dissociates in water to produce H^+ ions. **Base** is a substance that dissociates in water to produce OH^- ions.

A. The equilibrium system will move away from equilibrium when more acid is added.

B. The equilibrium system will move away from equilibrium when more base is added.

C. The equilibrium system will move away from equilibrium when more acid is added.

D. The equilibrium system will move away from equilibrium when more base is added.

Section 17.2

1. **Acid** is a substance that dissociates in water to produce H^+ ions. **Base** is a substance that dissociates in water to produce OH^- ions.

A. 1.000 mol

B. 1.000 mol

C. 0.500 mol

D. 0.500 mol

Section 17.3

1. The equilibrium system will move away from equilibrium when more acid is added.

A. 1.000 mol

B. 0.500 mol

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Ralph H. Petrucci, Robert K. Wismer

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