

Ch 17 Thermochemistry Practice Test**Matching**

Match each item with the correct statement below.

- | | |
|----------------|------------------|
| a. calorimeter | d. enthalpy |
| b. calorie | e. specific heat |
| c. joule | f. heat capacity |

- _____ 1. quantity of heat needed to raise the temperature of 1 g of water by 1°C
- _____ 2. SI unit of energy
- _____ 3. quantity of heat needed to change the temperature of 1 g of a substance by 1°C
- _____ 4. quantity of heat needed to change the temperature of an object by 1°C
- _____ 5. device used to measure the heat absorbed or released during a chemical or physical process
- _____ 6. heat content of a system at constant pressure

Match each item with the correct statement below.

- | | |
|---------------------------------|---------------------|
| a. heat of reaction | d. heat of fusion |
| b. heat of formation | e. heat of solution |
| c. Hess's law of heat summation | |

- _____ 7. the enthalpy change for a chemical reaction exactly as it is written
- _____ 8. the enthalpy change caused by dissolving a substance
- _____ 9. the energy required to melt a solid at its melting point
- _____ 10. the change in enthalpy that accompanies the formation of a compound from its elements
- _____ 11. states that if you add two or more thermochemical equations to give a final equation, you can also add the heats of reaction to give the final heat of reaction

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 12. What happens to the energy produced by burning gasoline in a car engine?
 - a. The energy is lost as heat in the exhaust.
 - b. The energy is transformed into work to move the car.
 - c. The energy heats the parts of the engine.
 - d. all of the above
- _____ 13. A piece of metal is heated, then submerged in cool water. Which statement below describes what happens?
 - a. The temperature of the metal will increase.
 - b. The temperature of the water will increase.
 - c. The temperature of the water will decrease.
 - d. The temperature of the water will increase and the temperature of the metal will decrease.
- _____ 14. How does a calorie compare to a joule?

a. A calorie is smaller than a joule.	c. A calorie is equal to a joule.
b. A calorie is larger than a joule.	d. The relationship cannot be determined.

17 Thermochemistry Matching

Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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Thermochemistry Models Applicable to a Vectorized Particle Simulation Brian Lee Haas, Stanford University. Department of Aeronautics and Astronautics, 1990 *Gas-phase Ion and Neutral Thermochemistry* Sharon G. Lias, 1988 Advances in Chemical Engineering Guy B. Marin, 2011-09-21 Understanding and modeling the kinetics of chemical reactions is crucial to any research and development effort aimed at process optimization and innovation This volume of *Advances in Chemical Engineering* provides four complementary points of view It reflects state of the art developments as well as views on the way to proceed by reporting on the efforts of a representative sample of research and development groups A first contribution by W H Green Jr sets the scene The author advocates a paradigm shift in chemical kinetics from postdictive to predictive models The contribution from the Politecnico di Milano reports on the tremendous experience accumulated over the years in the field of steam cracking one of the largest scale production processes of the petrochemical industry The Russian school of chemical kinetics is represented by a chapter on oxidation of alkanes this contribution addresses more philosophical issues The last chapter gives an indication of the state of the art in an industrial environment Provides original reviews Presents leading chemical engineers as authors Reviews state of the art developments *Technical Abstract Bulletin* , **Scientific and Technical Aerospace Reports** , 1977 *Dynamics of Explosions* A. L. Kuhl, 1988 Energy Research Abstracts , 1993

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