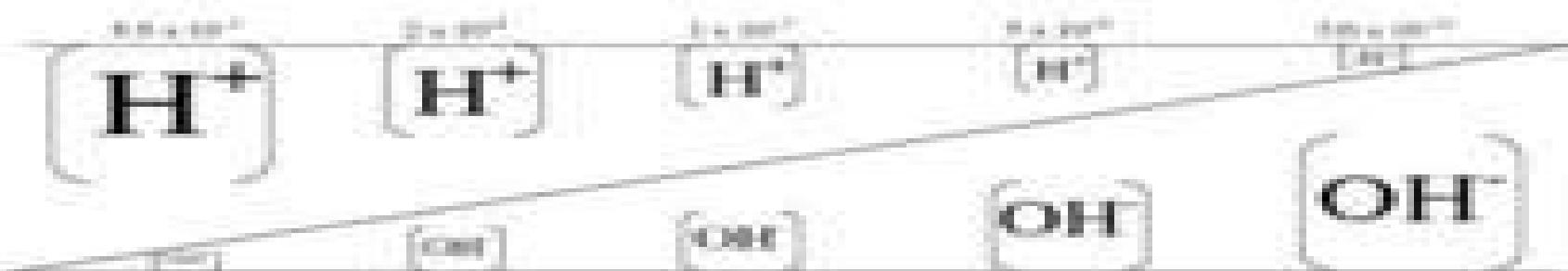


1.7 • Acid-Base Equilibrium

pH = -log₁₀[H⁺]

1. Write the pH of each solution given the hydronium ion concentration:

- 1.00 x 10⁻⁷ mol/L _____
- 1.00 x 10⁻¹⁰ mol/L _____
- 1.00 x 10⁻¹² mol/L _____
- 1.00 x 10⁻¹⁴ mol/L _____
- 1.00 x 10⁻¹⁶ mol/L _____
- 1.00 x 10⁻¹⁸ mol/L _____
- 1.00 x 10⁻²⁰ mol/L _____
- 1.00 x 10⁻²² mol/L _____
- 1.00 x 10⁻²⁴ mol/L _____
- 1.00 x 10⁻²⁶ mol/L _____
- 1.00 x 10⁻²⁸ mol/L _____
- 1.00 x 10⁻³⁰ mol/L _____
- 1.00 x 10⁻³² mol/L _____
- 1.00 x 10⁻³⁴ mol/L _____
- 1.00 x 10⁻³⁶ mol/L _____
- 1.00 x 10⁻³⁸ mol/L _____
- 1.00 x 10⁻⁴⁰ mol/L _____
- 1.00 x 10⁻⁴² mol/L _____
- 1.00 x 10⁻⁴⁴ mol/L _____
- 1.00 x 10⁻⁴⁶ mol/L _____
- 1.00 x 10⁻⁴⁸ mol/L _____
- 1.00 x 10⁻⁵⁰ mol/L _____
- 1.00 x 10⁻⁵² mol/L _____
- 1.00 x 10⁻⁵⁴ mol/L _____
- 1.00 x 10⁻⁵⁶ mol/L _____
- 1.00 x 10⁻⁵⁸ mol/L _____
- 1.00 x 10⁻⁶⁰ mol/L _____
- 1.00 x 10⁻⁶² mol/L _____
- 1.00 x 10⁻⁶⁴ mol/L _____
- 1.00 x 10⁻⁶⁶ mol/L _____
- 1.00 x 10⁻⁶⁸ mol/L _____
- 1.00 x 10⁻⁷⁰ mol/L _____
- 1.00 x 10⁻⁷² mol/L _____
- 1.00 x 10⁻⁷⁴ mol/L _____
- 1.00 x 10⁻⁷⁶ mol/L _____
- 1.00 x 10⁻⁷⁸ mol/L _____
- 1.00 x 10⁻⁸⁰ mol/L _____
- 1.00 x 10⁻⁸² mol/L _____
- 1.00 x 10⁻⁸⁴ mol/L _____
- 1.00 x 10⁻⁸⁶ mol/L _____
- 1.00 x 10⁻⁸⁸ mol/L _____
- 1.00 x 10⁻⁹⁰ mol/L _____
- 1.00 x 10⁻⁹² mol/L _____
- 1.00 x 10⁻⁹⁴ mol/L _____
- 1.00 x 10⁻⁹⁶ mol/L _____
- 1.00 x 10⁻⁹⁸ mol/L _____
- 1.00 x 10⁻¹⁰⁰ mol/L _____

2. Calculate the pH and pOH of each solution given the hydronium ion concentration:

- 1.00 x 10⁻⁷ mol/L _____
- 1.00 x 10⁻¹⁰ mol/L _____
- 1.00 x 10⁻¹² mol/L _____
- 1.00 x 10⁻¹⁴ mol/L _____
- 1.00 x 10⁻¹⁶ mol/L _____
- 1.00 x 10⁻¹⁸ mol/L _____
- 1.00 x 10⁻²⁰ mol/L _____
- 1.00 x 10⁻²² mol/L _____
- 1.00 x 10⁻²⁴ mol/L _____
- 1.00 x 10⁻²⁶ mol/L _____
- 1.00 x 10⁻²⁸ mol/L _____
- 1.00 x 10⁻³⁰ mol/L _____
- 1.00 x 10⁻³² mol/L _____
- 1.00 x 10⁻³⁴ mol/L _____
- 1.00 x 10⁻³⁶ mol/L _____
- 1.00 x 10⁻³⁸ mol/L _____
- 1.00 x 10⁻⁴⁰ mol/L _____
- 1.00 x 10⁻⁴² mol/L _____
- 1.00 x 10⁻⁴⁴ mol/L _____
- 1.00 x 10⁻⁴⁶ mol/L _____
- 1.00 x 10⁻⁴⁸ mol/L _____
- 1.00 x 10⁻⁵⁰ mol/L _____
- 1.00 x 10⁻⁵² mol/L _____
- 1.00 x 10⁻⁵⁴ mol/L _____
- 1.00 x 10⁻⁵⁶ mol/L _____
- 1.00 x 10⁻⁵⁸ mol/L _____
- 1.00 x 10⁻⁶⁰ mol/L _____
- 1.00 x 10⁻⁶² mol/L _____
- 1.00 x 10⁻⁶⁴ mol/L _____
- 1.00 x 10⁻⁶⁶ mol/L _____
- 1.00 x 10⁻⁶⁸ mol/L _____
- 1.00 x 10⁻⁷⁰ mol/L _____
- 1.00 x 10⁻⁷² mol/L _____
- 1.00 x 10⁻⁷⁴ mol/L _____
- 1.00 x 10⁻⁷⁶ mol/L _____
- 1.00 x 10⁻⁷⁸ mol/L _____
- 1.00 x 10⁻⁸⁰ mol/L _____
- 1.00 x 10⁻⁸² mol/L _____
- 1.00 x 10⁻⁸⁴ mol/L _____
- 1.00 x 10⁻⁸⁶ mol/L _____
- 1.00 x 10⁻⁸⁸ mol/L _____
- 1.00 x 10⁻⁹⁰ mol/L _____
- 1.00 x 10⁻⁹² mol/L _____
- 1.00 x 10⁻⁹⁴ mol/L _____
- 1.00 x 10⁻⁹⁶ mol/L _____
- 1.00 x 10⁻⁹⁸ mol/L _____
- 1.00 x 10⁻¹⁰⁰ mol/L _____

3. Calculate the pH and pOH of each solution given the hydronium ion concentration:

- 1.00 x 10⁻⁷ mol/L _____
- 1.00 x 10⁻¹⁰ mol/L _____
- 1.00 x 10⁻¹² mol/L _____
- 1.00 x 10⁻¹⁴ mol/L _____
- 1.00 x 10⁻¹⁶ mol/L _____
- 1.00 x 10⁻¹⁸ mol/L _____
- 1.00 x 10⁻²⁰ mol/L _____
- 1.00 x 10⁻²² mol/L _____
- 1.00 x 10⁻²⁴ mol/L _____
- 1.00 x 10⁻²⁶ mol/L _____
- 1.00 x 10⁻²⁸ mol/L _____
- 1.00 x 10⁻³⁰ mol/L _____
- 1.00 x 10⁻³² mol/L _____
- 1.00 x 10⁻³⁴ mol/L _____
- 1.00 x 10⁻³⁶ mol/L _____
- 1.00 x 10⁻³⁸ mol/L _____
- 1.00 x 10⁻⁴⁰ mol/L _____
- 1.00 x 10⁻⁴² mol/L _____
- 1.00 x 10⁻⁴⁴ mol/L _____
- 1.00 x 10⁻⁴⁶ mol/L _____
- 1.00 x 10⁻⁴⁸ mol/L _____
- 1.00 x 10⁻⁵⁰ mol/L _____
- 1.00 x 10⁻⁵² mol/L _____
- 1.00 x 10⁻⁵⁴ mol/L _____
- 1.00 x 10⁻⁵⁶ mol/L _____
- 1.00 x 10⁻⁵⁸ mol/L _____
- 1.00 x 10⁻⁶⁰ mol/L _____
- 1.00 x 10⁻⁶² mol/L _____
- 1.00 x 10⁻⁶⁴ mol/L _____
- 1.00 x 10⁻⁶⁶ mol/L _____
- 1.00 x 10⁻⁶⁸ mol/L _____
- 1.00 x 10⁻⁷⁰ mol/L _____
- 1.00 x 10⁻⁷² mol/L _____
- 1.00 x 10⁻⁷⁴ mol/L _____
- 1.00 x 10⁻⁷⁶ mol/L _____
- 1.00 x 10⁻⁷⁸ mol/L _____
- 1.00 x 10⁻⁸⁰ mol/L _____
- 1.00 x 10⁻⁸² mol/L _____
- 1.00 x 10⁻⁸⁴ mol/L _____
- 1.00 x 10⁻⁸⁶ mol/L _____
- 1.00 x 10⁻⁸⁸ mol/L _____
- 1.00 x 10⁻⁹⁰ mol/L _____
- 1.00 x 10⁻⁹² mol/L _____
- 1.00 x 10⁻⁹⁴ mol/L _____
- 1.00 x 10⁻⁹⁶ mol/L _____
- 1.00 x 10⁻⁹⁸ mol/L _____
- 1.00 x 10⁻¹⁰⁰ mol/L _____

4. pH levels of various acids and bases in common acids _____ usually in either the standard pH scale _____

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