

**15.2****HOMOGENEOUS AQUEOUS SYSTEMS****Section Review****Objectives**

- Distinguish between a solvent and a solute
- Describe what happens in the solution process
- Explain why all ionic compounds are electrolytes
- Demonstrate how to write the formula for a hydrate

**Vocabulary**

- |                    |                  |                      |
|--------------------|------------------|----------------------|
| • aqueous solution | • solvation      | • strong electrolyte |
| • solvent          | • electrolyte    | • weak electrolyte   |
| • solute           | • nonelectrolyte | • hydrate            |

**Key Equation**

- $\text{Percent H}_2\text{O} = \frac{\text{mass of water}}{\text{mass of hydrate}} \times 100\%$

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water is a polar liquid and an excellent 1 for many substances. Aqueous solutions are 2 mixtures of ions or molecules in water. The solubility of a solute depends on solute-solvent interactions. A good rule to remember is 3.

Substances that dissolve as ions are known as 4. A solute that is completely ionized in solution is a 5 electrolyte. A weak electrolyte is only 6 ionized. A solution of an electrolyte will 7 an electric current, whereas a solution of a 8 is nonconducting.

Many crystals are 9; they contain water of hydration. In the process called 10, the water of hydration is lost from a hydrate that is exposed to air.

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## 15 2 Homogeneous Aqueous Systems Pg 379 Answers

**N Noddings**



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