

## Periodic Trends Worksheet

Use the periodic table and your knowledge of periodic trends to answer the following questions.

1. Which atom in each pair has the larger atomic radius?  
a) Li or **K**      b) **Ca** or Na      c) **Ga** or B      d) O or **C**      e) Cl or **Br**  
f) Be or **Ba**      g) **Si** or S      h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? from left to right in a period?

Top to bottom: increases due to adding energy levels   Left to Right: decreases

3. Why do atoms get smaller as you move left to right in a period?  
due to increased nuclear charge and no added shielding

4. Which element in each pair has a larger ionization energy?  
a) Na or **O**      b) **Be** or Ba      c) **Ar** or F      d) **Cu** or Ra      e) I or **Ne**  
f) K or **V**      g) **Ca** or Fr      h) W or **Se**

5. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.  
Positive ions are smaller than the neutral atom it comes from, negative ions are larger than the neutral atom it comes from

6. Which particle has the larger radius in each atom/ion pair?  
a) Na,  $\text{Na}^+$       b) S,  $\text{S}^{2-}$       c) I,  $\text{I}^-$       d) **Al**,  $\text{Al}^{3+}$

7. What is ionization energy? What is first ionization energy?

The amount of energy needed to remove an electron from an atom. First ionization energy is the energy required to remove a single valence electron from an atom.

8. What is the periodic trend for first ionization energy?

It increases going from left to right due to increased nuclear charge, and decreases going down a group due to increased distance from the nucleus and increased shielding.

9. Arrange the following groups of elements in order of increasing ionization energy.  
a) Be, Mg, Sr      b) Bi, Cs, Ba      c) Na, Al, S  
high, middle, low      high, low, middle      low, middle, high

10. Which element in each pair has a higher electronegativity value?  
a) Cl, **F**      b) C, **N**      c) **Mg**, Ne      d) As, **Ca**

11. What is the periodic trend for electronegativity? Increases going left to right, decreases going down.

# 15 Periodic Trends Answers

**Charles H. Corwin**



## 15 Periodic Trends Answers:

*Elements and the Periodic Table, Grades 5 - 12* Theodore A. Abbgy, 2001-01-01 This informative classroom supplement is a great introduction to the periodic table explored in sequential form It includes activities transparency masters a teacher s guide an element game quizzes tests rubrics and answer keys Unit topics include discovering what elements are the uses of the elements element symbols periodic table organization and more Mark Twain Media Publishing Company specializes in providing captivating supplemental books and decorative resources to complement middle and upper grade classrooms Designed by leading educators the product line covers a range of subjects including mathematics sciences language arts social studies history government fine arts and character Mark Twain Media also provides innovative classroom solutions for bulletin boards and interactive whiteboards Since 1977 Mark Twain Media has remained a reliable source for a wide variety of engaging classroom resources

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**Science For Tenth Class Part 2 Chemistry** Lakhmir Singh & Manjit Kaur, A series of six books for Classes IX and X according to the CBSE syllabus Each class divided into 3 parts Part 1 Physics Part 2 Chemistry Part 3 Biology **Arihant CBSE Science Term 2 Class 10 for 2022 Exam (Cover Theory and MCQs)** Naman jain, Jyoti Agarwal, Anam Aarzo, 2021-11-20 With newly introduced 2 Term Examination Pattern CBSE has eased out the pressure of preparation of subjects and cope up with lengthy syllabus Introducing Arihant s CBSE TERM II 2022 Series the first of its kind that gives complete emphasize on the rationalize syllabus of Class 10th 12th The all new CBSE Term II 2022 Science of Class 10th provides explanation and guidance to the syllabus required to study efficiently and succeed in the exams The book provides topical coverage of all the chapters in a complete and comprehensive manner Covering the 50% of syllabus as per Latest Term wise pattern 2021 22 this book consists of 1 Complete Theory in each Chapter covering all topics 2 Case Based Short and Long Answer Type Question in each chapter 3 Coverage of NCERT NCERT Exemplar Board Exams Questions 4 Complete and Detailed explanations for each question 5 3 Practice papers base on entire Term II Syllabus Table of Content Carbon and its compounds Periodic Classification of Elements How do Organisms Reproduce Heredity and Evolution Electricity Magnetic Effects and Electric Current Our Environment Practice Paper 1 3 *Prentice Hall Chemistry* , 2000 Science Games and Puzzles, Grades 5 - 8 Schyrlet Cameron, Carolyn Craig, 2012-01-03 This book promotes science vocabulary building increases student readability levels and facilitates concept development through fun and challenging puzzles games and activities

**Chemical Interactions** , 2005 **Campbell's Accent on Numbers Quiz Book** John P. Campbell, 2000-10

Fundamentals of General, Organic, and Biological Chemistry John R. Holum, 1982 User friendly this comprehensive and up to date edition is organized according to the principle that life in both health and disease has a molecular basis Carefully

worked example exercises enhanced by Analysis and Check sections enable students to examine and think about a problem not just answer it Includes numerous environmental applications      **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

**A Complete Preparation for the MCAT.: Knowledge & comprehension of science** James L. Flowers,1988

Chemistry, Study Guide Bernice G. Segal,Peter S. Shenkin,1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460      *Holt Chemistry* Ralph Thomas Myers,2004      Science in Your World: Teacher resource masters Jay K.

Hackett,1991      Programmed Topics in General Chemistry Armine D. Paul,1971      **Chemistry** Theodore Lawrence Brown,1988      **Question and Answer Encyclopedia** John Farndon,1998      **Introductory Chemistry** Charles H. Corwin,2005 For one semester courses in Basic Chemistry Introduction to Chemistry and Preparatory Chemistry and the first term of Allied Health Chemistry This text is carefully crafted to help students learn chemical skills and concepts more effectively Corwin covers math and problem solving early in the text he builds student confidence and skills through innovative problem solving pedagogy and technology formulated to meet student needs      Nursing School Entrance Exam Review William H. Brown,Frederick A. Bettelheim,Jerry March,2004-04

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web abstract in this experiment we measured  $g$  by measuring the period of a pendulum of a known length we measured  $g = 7.65 \pm 0.378 \text{ m/s}^2$   $g = 7.65 \pm 0.378 \text{ m/s}^2$  this corresponds to a relative difference of 22.22% with the accepted value  $9.8 \text{ m/s}^2$  and our result is not consistent with the accepted value

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