

15.2**HOMOGENEOUS AQUEOUS SYSTEMS****Section Review****Objectives**

- Distinguish between a solvent and a solute
- Describe what happens in the solution process
- Explain why all ionic compounds are electrolytes
- Demonstrate how to write the formula for a hydrate

Vocabulary

• aqueous solution	• solvation	• strong electrolyte
• solvent	• electrolyte	• weak electrolyte
• solute	• nonelectrolyte	• hydrate

Key Equation

$$\bullet \text{ Percent H}_2\text{O} = \frac{\text{mass of water}}{\text{mass of hydrate}} \times 100\%$$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water is a polar liquid and an excellent 1 for many substances. Aqueous solutions are 2 mixtures of ions or molecules in water. The solubility of a solute depends on solute-solvent interactions. A good rule to remember is 3.

Substances that dissolve as ions are known as 4. A solute that is completely ionized in solution is a 5 electrolyte. A weak electrolyte is only 6 ionized. A solution of an electrolyte will 7 an electric current, whereas a solution of a 8 is nonconducting.

Many crystals are 9; they contain water of hydration. In the process called 10, the water of hydration is lost from a hydrate that is exposed to air.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

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the principles response behavior ion selectivity and detection limits of solid state membrane electrodes This part also examines several important extensions and modifications of the Sandblom Eisenman Walker theory the characteristics of neutral carrier membrane electrodes and the theory of glass electrodes Monthly Bibliography of Medical Reviews ,1972

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