

**14-1 Review and Reinforcement (continued)**

Complete each of the sentences below.

15. The stronger the intermolecular forces in a liquid, the \_\_\_\_\_ its boiling point.
16. The state of a substance at room temperature depends on \_\_\_\_\_.
17. Water molecules in an ice cube are held together by \_\_\_\_\_ forces or, more specifically, \_\_\_\_\_ bonds.

On the line at the left, write the letter of the term that matches each description below. Each choice can be used once, more than once, or not at all.

a. dispersion      b. dipole-dipole      c. hydrogen bonding

18. the force in between  $\text{SO}_3$  molecules
19. the force that accounts for  $\text{HF}$  being a liquid while  $\text{H}_2$  and  $\text{F}_2$  are gases
20. the force that depends on temporary, induced dipoles
21. the only intermolecular force that exists in noble gases
22. the attractive force between neighboring dipoles
23. the force that arises because of the large difference in electronegativities in the N-H bonds

Answer the following question as directed.

24. Which of the following statements could be used to explain why water has an unusually high boiling point? Make a checkmark on the line for all that apply.

\_\_\_\_\_ a. Water molecules have a strong intermolecular force.  
\_\_\_\_\_ b. There are no nitrogen atoms in water.  
\_\_\_\_\_ c. Oxygen is more electronegative than hydrogen.  
\_\_\_\_\_ d. Water is an angular molecule.  
\_\_\_\_\_ e. Oxygen atoms are smaller than hydrogen atoms.  
\_\_\_\_\_ f. Water does not contain metallic bonds.  
\_\_\_\_\_ g. There is a large difference in electronegativity between oxygen and hydrogen.

## 14 Review And Reinforcement

**Lingjun Ying**

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## **Table of Contents 14 Review And Reinforcement**

1. Understanding the eBook 14 Review And Reinforcement
  - The Rise of Digital Reading 14 Review And Reinforcement
  - Advantages of eBooks Over Traditional Books
2. Identifying 14 Review And Reinforcement
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 14 Review And Reinforcement
  - User-Friendly Interface
4. Exploring eBook Recommendations from 14 Review And Reinforcement
  - Personalized Recommendations
  - 14 Review And Reinforcement User Reviews and Ratings
  - 14 Review And Reinforcement and Bestseller Lists
5. Accessing 14 Review And Reinforcement Free and Paid eBooks

- 14 Review And Reinforcement Public Domain eBooks
- 14 Review And Reinforcement eBook Subscription Services
- 14 Review And Reinforcement Budget-Friendly Options

6. Navigating 14 Review And Reinforcement eBook Formats

- ePUB, PDF, MOBI, and More
- 14 Review And Reinforcement Compatibility with Devices
- 14 Review And Reinforcement Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 14 Review And Reinforcement
- Highlighting and Note-Taking 14 Review And Reinforcement
- Interactive Elements 14 Review And Reinforcement

8. Staying Engaged with 14 Review And Reinforcement

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 14 Review And Reinforcement

9. Balancing eBooks and Physical Books 14 Review And Reinforcement

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 14 Review And Reinforcement

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 14 Review And Reinforcement

- Setting Reading Goals 14 Review And Reinforcement
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 14 Review And Reinforcement

- Fact-Checking eBook Content of 14 Review And Reinforcement
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

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- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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