

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PV=nRT", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems.

If pressure is needed in kPa then convert by multiplying by $101.3kPa / 1atm$ to get $R = 8.31 \frac{kPa \cdot L}{K \cdot mole}$

1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$
$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 3.1 liters, and a temperature of 87°C , how many moles of gas do I have?

$$PV = nRT$$
$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 3.1 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.1586 \text{ mol}$$

3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$
$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

or

$$P = 166.24 \text{ kPa}$$

4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56°C , what is the volume of the container that the gas is in?

$$PV = nRT$$
$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 329 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

5) If I have 17 moles of gas at a temperature of 67°C , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$
$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

or

$$P = 5.34 \text{ atm}$$

$$P = 590.61 \text{ kPa}$$

6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$
$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

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RS Peters

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