

Standardized Test prep Answers

Chapter 1 Page 25			Chapter 2 Page 63		
1. C	4. B	7. C	1. B	4. B	7. A
2. C	5. C	8. A	2. C	5. A	8. B
3. D	6. C	9. B	3. C	6. B	9. B
10. Answers may vary.			10. 2 cups; direct proportion		
11. Observations C and G are chemical properties; the others are physical properties.			11. Statements A and B contain exact numbers.		
12. In a chemical change, one or more substances are converted into different substances. A physical change does not involve a change in the identity of the substance or substances present.			12. The type of fertilizer is the variable being tested. Control factors are the types of radishes, the amount of water and the amount of sunshine. One control row should be planted under the same control factors but with no fertilizer. There are at least four things that could be used to determine the results: size, quantity, appearance, and taste. Analysis might include bar graphs of each of these measurements for each of the five fertilizer types and the no-fertilizer control row.		
13. metals: shiny; good conductors of heat; good conductors of electricity; malleable or ductile; most are solids at room temperature; nonmetals: poor conductors of heat; poor conductors of electricity; many are gases at room temperature; those that are solids are brittle rather than malleable or ductile; metalloids: properties intermediate between those of metals and nonmetals; less malleable than metals but not as brittle as solid nonmetals; most are semiconductors of electricity			13. A unit must be defined in a way that does not depend on the circumstances of the measurement. Not every thumbnail is the same size.		
Chapter 3 Page 93			Chap 4 Page 129		
1. C	4. A	7. B	1. A	4. D	7. C
2. D	5. B	8. D	2. D	5. B	8. B
3. C	6. A	9. C	3. A	6. B	9. C
10. Argon-40 has 22 neutrons ($40 - 18 = 22$), and potassium-40 has 21 neutrons ($40 - 19 = 21$).			10. The color will be orange. Converting energy into frequency gives 4.8×10^{14} , which corresponds to the frequency of orange light.		
11. 6.17 g			11. 2.4×10^{-1}		
12. All cathode rays are the same, regardless of their source. Therefore, the particles responsible for the cathode rays must be present in all atoms. The particles are electrons.			12. Electrons in atoms can occupy orbitals of only specific energies. When an atom is excited, the electron is no longer in the ground state. When the electron returns to a lower energy level, light is emitted. Because only specific energies are allowed, certain wavelengths of light are emitted, giving rise to the individual lines in the spectrum.		
13. When the average atomic mass is calculated, it is 10.811. Because the atomic mass is the same as the atomic mass of boron, mythium was not a new element.			13. Photons of blue light are higher energy than photons of yellow light. Electrons can be emitted only when a photon of sufficient energy strikes the surface of the metal. Therefore, the energy of blue light is greater than the threshold energy, but the energy of yellow light is not.		
Chapter 5 Page 1			Chapter 6 Page 215		
1. C	2. C	7. A	1. A	2. C	7. A
2. A	3. A	8. B	3. B	4. D	8. B
3. A	4. A	9. C	5. A	6. D	9. C
10. S: 2251 kJ/mol; Cl: 2297 kJ/mol; Ar: 2666 kJ/mol; K: 3053 kJ/mol; Ca: 1349 kJ/mol; Sc: 1235 kJ/mol; Ti: 1310 kJ/mol. For the second ionization, the general trend is for increasing IE2 across the period in Groups 2–18 with Group 1 having the highest IE2. IE2 decreases going down a group.			10. Hybridization explains how the orbitals of an atom become rearranged when the atom forms covalent bonds.		
11. Group 16 most commonly forms 2- ions, because these elements require only two more electrons to fill their shell (obtain a noble-gas configuration).			11. Ionic crystals are brittle because shifting of the layers of ions results in large repulsive forces that cause the layers to part completely.		
12. Naphthalene has the stronger intermolecular forces even though it is non-polar, because its boiling point is higher than that of acetic acid. Boiling point is directly correlated to strength of intermolecular forces; the stronger the intermolecular forces, the more energy needed to break all the intermolecular forces, and therefore the higher the boiling point. Naphthalene is so large that its dispersion forces are greater than the sum of the dispersion forces and hydrogen bonding in acetic acid.			12. Naphthalene has the stronger intermolecular forces even though it is non-polar, because its boiling point is higher than that of acetic acid. Boiling point is directly correlated to strength of intermolecular forces; the stronger the intermolecular forces, the more energy needed to break all the intermolecular forces, and therefore the higher the boiling point. Naphthalene is so large that its dispersion forces are greater than the sum of the dispersion forces and hydrogen bonding in acetic acid.		
13. There is a repulsive region at small bond distances, a minimum well at an intermediate bond distance, and an asymptotic region at large bond distances. The repulsive region is due to the repulsion between the two positive nuclei at distances where the nuclei are close to each other. The asymptotic region is where the atoms are far apart and therefore do not interact with each other. The potential energy decreases as the atoms come closer together and are attracted to each other. The minimum well is the region where the attractive forces are balanced by the repulsive forces. This distance is the equilibrium bond length.			13. There is a repulsive region at small bond distances, a minimum well at an intermediate bond distance, and an asymptotic region at large bond distances. The repulsive region is due to the repulsion between the two positive nuclei at distances where the nuclei are close to each other. The asymptotic region is where the atoms are far apart and therefore do not interact with each other. The potential energy decreases as the atoms come closer together and are attracted to each other. The minimum well is the region where the attractive forces are balanced by the repulsive forces. This distance is the equilibrium bond length.		

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Table of Contents 10 6 Standardized Test Prep Key

1. Understanding the eBook 10 6 Standardized Test Prep Key
 - The Rise of Digital Reading 10 6 Standardized Test Prep Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 10 6 Standardized Test Prep Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 10 6 Standardized Test Prep Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 10 6 Standardized Test Prep Key
 - Personalized Recommendations
 - 10 6 Standardized Test Prep Key User Reviews and Ratings
 - 10 6 Standardized Test Prep Key and Bestseller Lists
5. Accessing 10 6 Standardized Test Prep Key Free and Paid eBooks
 - 10 6 Standardized Test Prep Key Public Domain eBooks
 - 10 6 Standardized Test Prep Key eBook Subscription Services
 - 10 6 Standardized Test Prep Key Budget-Friendly Options

6. Navigating 10 6 Standardized Test Prep Key eBook Formats
 - ePUB, PDF, MOBI, and More
 - 10 6 Standardized Test Prep Key Compatibility with Devices
 - 10 6 Standardized Test Prep Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 10 6 Standardized Test Prep Key
 - Highlighting and Note-Taking 10 6 Standardized Test Prep Key
 - Interactive Elements 10 6 Standardized Test Prep Key
8. Staying Engaged with 10 6 Standardized Test Prep Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 10 6 Standardized Test Prep Key
9. Balancing eBooks and Physical Books 10 6 Standardized Test Prep Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 10 6 Standardized Test Prep Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 10 6 Standardized Test Prep Key
 - Setting Reading Goals 10 6 Standardized Test Prep Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 10 6 Standardized Test Prep Key
 - Fact-Checking eBook Content of 10 6 Standardized Test Prep Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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