

# Percent Composition

## Practice & Worksheet | Chemistry Guided & Independent Practice

### PERCENT Composition

Find the percentage of elements with the compound.



Cu

S

O

H



Na

C

O

H

### PERCENT Composition



Percent composition is a fundamental concept in chemistry that expresses the relative masses of each element in a compound. It is calculated by dividing the mass of each element by the molar mass of the compound and converting it into a percentage.

Formula for Calculation:

$$\text{Percentage} = \left( \frac{\text{Mass of element in the compound}}{\text{Molar mass of the compound}} \right) \times 100$$

Example Calculation:

1. Find the Percent Composition of  $\text{H}_2\text{O}$  (Water)

- Determine the molar mass of  $\text{H}_2\text{O}$
- Hydrogen (2):  $1 \text{ g/mol} \times 2 = 2 \text{ g/mol}$
- Oxygen (16):  $16 \text{ g/mol} \times 1 = 16 \text{ g/mol}$
- Total molar mass = 18 g/mol

2. Calculate the percent composition of each element:

- Percent Hydrogen (H):

$$\frac{2}{18} \times 100 = 11.11\%$$

- Percent Oxygen (O):

$$\frac{16}{18} \times 100 = 88.89\%$$

Thus, the percent composition of  $\text{H}_2\text{O}$  is:

IT

1. Find the Percent Composition of the following:

1. $\text{H}_2\text{O}$	Mass = 1800	
2. $\text{H}_2\text{SO}_4$	Mass = 900	
3. $\text{NaCl}$	Mass = 58.5	
4. $\text{CO}_2$	Mass = 44	
5. $\text{CH}_4$	Mass = 16	
6. $\text{C}_2\text{H}_6$	Mass = 30	
7. $\text{C}_3\text{H}_8$	Mass = 44	
8. $\text{C}_4\text{H}_{10}$	Mass = 58	
9. $\text{C}_5\text{H}_{12}$	Mass = 72	
10. $\text{C}_6\text{H}_{14}$	Mass = 98	
11. $\text{C}_7\text{H}_{16}$	Mass = 112	
12. $\text{C}_8\text{H}_{18}$	Mass = 126	
13. $\text{C}_9\text{H}_{20}$	Mass = 140	
14. $\text{C}_{10}\text{H}_{22}$	Mass = 154	

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**Singhal Atul**



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