

Percent Composition

Practice & Worksheet | Chemistry Guided & Independent Practice

PERCENT Composition



Cu

S

O

H



Na

C

O

H

PERCENT Composition



Percent composition is a fundamental concept in chemistry that represents the relative measure of each element in a compound. This is calculated by dividing the mass of each element by the total mass of the compound and converting it into a percentage.

Formula for Calculation

$$\text{Percentage} = \left(\frac{\text{Mass of element in the compound}}{\text{Total mass of the compound}} \right) \times 100$$

Example Calculations

- Find the Percent Composition of H_2O_2 .
 - Determine the molar mass of H_2O_2 .
 - Hydrogen (H) is given as $1 \times 2 = 2$ g/mole
 - Oxygen (O) is given as $16 \times 2 = 32$ g/mole
 - Total molar mass = 34.00 g/mole
 - Calculate the percent composition of each element.
 - Percent hydrogen (H):
$$\frac{2}{34} \times 100\% = 5.88\%$$
 - Percent oxygen (O):
$$\frac{32}{34} \times 100\% = 94.12\%$$

Thus, the percent composition of H_2O_2 :

IT

100

100% of each element in the following

Element	Mass	Percent
Cu	64.00	35.5%
S	32.00	15.0%
O	16.00	7.2%
H	1.00	0.5%

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Cu	64.00	35.5%
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Worksheet

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Singhal Atul

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