

CHEMISTRY 11 – WORD EQUATIONS AND TYPES OF CHEMICAL REACTIONS
WORKSHEET

- 1) Translate the following word equations into chemical symbols and balance the resulting equations. Do not include the phases.

- (a) Potassium + water \rightarrow potassium hydroxide + hydrogen



- (b) Nitrogen + hydrogen \rightarrow ammonia



- 2) In each of the following write a balanced chemical equation, including the phases.

- (a) Liquid phosphoric acid reacts with aqueous barium hydroxide to give water and a precipitate of barium phosphate



- (b) Solid aluminium oxide and aqueous sulphuric acid produce water and aqueous aluminium sulphate



- (c) Nitrogen trifluoride gas and hydrogen gas react to form nitrogen gas and gaseous hydrogen fluoride



- (d) Powdered sodium carbonate and aqueous hydrobromic acid react to form carbon dioxide gas, aqueous sodium bromide and water



- 3) Classify the following reactions.

Reaction	Type
$2 \text{C}_2\text{H}_2 + 5 \text{O}_2 \rightarrow 4 \text{CO}_2 + 2 \text{H}_2\text{O}$	Combustion
$\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$	Single Replacement
$4 \text{Na} + \text{O}_2 \rightarrow 2 \text{Na}_2\text{O}$	Synthesis
$\text{Sn}(\text{OH})_4 + 4 \text{HBr} \rightarrow 4 \text{H}_2\text{O} + \text{SnBr}_4$	Neutralization
$2 \text{Fe}(\text{NO}_3)_3 + 3 \text{MgS} \rightarrow \text{Fe}_2\text{S}_3 + 3 \text{Mg}(\text{NO}_3)_2$	Double Replacement
$2 \text{N}_2\text{O} \rightarrow 2 \text{N}_2 + \text{O}_2$	Decomposition

10 1 Reactions And Equations Answers

Theodore Weicker

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