

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|--|---|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 ²³ | |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition | |
| _____ the lowest whole number ratio of the atoms of the elements in a compound _____ empirical formula | |
| _____ the volume of 1 mole of any gas at STP _____ 22.4 L | |
| _____ the smallest particle of a compound is a compound _____ molar mass | |
| _____ the number representing 6.02 $\times 10^{23}$ is Avogadro's number particles of a substance _____ molar | |
| _____ 1 mole = 6.02 $\times 10^{23}$ particles _____ empirical formula | |
| _____ the number representing the number of atoms of an element in a compound (2H ₂ O _____ 2 hydrogen atoms, 1 oxygen atom) (2H ₂ O _____ 4 H, 2 O) | |

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Completion Use the completion strategy to strengthen knowledge of the terms and your understanding of the concepts involved with the terms of the definitions completed with a term (don't guess) or explain.

Measuring whole units (counting, weighing, or measuring volume for a single quantity) called the _____ mass _____ . The number of representative particles in a large (g) substance is _____ Avogadro's number _____ .

6.02 $\times 10^{23}$ is the number of particles of a compound, substance, or mixture that _____ molar mass _____ of the substance (g/mol) is calculated. Other two techniques for measuring the mass (m) of a substance _____ molar mass _____ of the compound. These are _____ (g/mol) or _____ (g/mol) _____ representative particles the mass of one substance.

True or False

_____ 1 mole of any substance contains 6.02 $\times 10^{23}$ particles.

_____ 1 mole of a substance contains 6.02 $\times 10^{23}$ atoms.

_____ 1 mole of a substance contains 6.02 $\times 10^{23}$ particles of the substance. **True**

_____ 1 mole of H₂O is composed of one mole of hydrogen and one mole of oxygen.

_____ 1 mole of a substance contains 6.02 $\times 10^{23}$ particles. **True**

Problems

10 Chemical Quantities Chapter Quiz Answer Key

Martha Mackin



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