



IGCSE A* MODEL ANSWER

0620 CHEMISTRY

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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.



Table 3.1 shows some bond energies.

Table 3.1

Bond	N≡N	H-H
Bond energy in kJ/mol	946	435

Use the bond energies in Table 3.1 and 3.2 to calculate the bond energy of an H-Cl bond in HCl.

Use the following steps:

- Calculate the energy needed to break bonds in the reactants.

$$(946) + 1(435)$$

$$946 + 1(435)$$

1381

- Calculate the energy released when bonds form in the products.

$$\text{heat released} = \text{heat taken} = -191$$

$$1381 = 35 + -191$$
$$1381 = 35 + -191$$

-1560

- Calculate the energy of an H-Cl bond.

$$1/2 (191) = 95.5$$

$$1/2 = 95.5$$

95.5

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Clemens Wendtner

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Table of Contents 0620 31 M J 14 Mark Scheme

1. Understanding the eBook 0620 31 M J 14 Mark Scheme
 - The Rise of Digital Reading 0620 31 M J 14 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying 0620 31 M J 14 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 0620 31 M J 14 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from 0620 31 M J 14 Mark Scheme
 - Personalized Recommendations
 - 0620 31 M J 14 Mark Scheme User Reviews and Ratings
 - 0620 31 M J 14 Mark Scheme and Bestseller Lists
5. Accessing 0620 31 M J 14 Mark Scheme Free and Paid eBooks
 - 0620 31 M J 14 Mark Scheme Public Domain eBooks

- 0620 31 M J 14 Mark Scheme eBook Subscription Services
- 0620 31 M J 14 Mark Scheme Budget-Friendly Options

6. Navigating 0620 31 M J 14 Mark Scheme eBook Formats

- ePub, PDF, MOBI, and More
- 0620 31 M J 14 Mark Scheme Compatibility with Devices
- 0620 31 M J 14 Mark Scheme Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 0620 31 M J 14 Mark Scheme
- Highlighting and Note-Taking 0620 31 M J 14 Mark Scheme
- Interactive Elements 0620 31 M J 14 Mark Scheme

8. Staying Engaged with 0620 31 M J 14 Mark Scheme

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 0620 31 M J 14 Mark Scheme

9. Balancing eBooks and Physical Books 0620 31 M J 14 Mark Scheme

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 0620 31 M J 14 Mark Scheme

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 0620 31 M J 14 Mark Scheme

- Setting Reading Goals 0620 31 M J 14 Mark Scheme
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 0620 31 M J 14 Mark Scheme

- Fact-Checking eBook Content of 0620 31 M J 14 Mark Scheme
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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