

## Experiment 6: Activity Series: Single Replacement Reactions

**Required reading:** Ebbing, 11<sup>th</sup> Edition Chapter 4.

- Molecular and Ionic equations
- Displacement reactions
- Balancing redox equations

### Learning Goals:

- To observe reactions of metals with acids and salt solutions using activity series of metals.
- To be able to write and balance single replacement reactions using activity series of metals.
- To be able to write ionic and net ionic equations for the above reactions.

### Background information and theory:

One of the most important characteristics of a metal is its activity (reactivity). The activity of a metal is its ability to react with nonmetals. In such reaction, a metal typically loses electrons to the nonmetal to form cations and the product is an ionic compound. Different metals have different activity. Some metals are very active such as lithium, sodium, potassium, and cesium; some are slowly active such as tin, lead, and iron; and some are not active at all (inactive) such as copper, silver, and mercury. Inactive metals are those which resist oxidation. These are often called the noble metals.

An **activity series** is the ranking of metals according to their reactivity. In an activity series metals are arranged in order of decreasing ability to lose electrons. Generally, the more reactive metals are placed at the left of the list.

**Li > K > Ba > Ca > Na > Mg > Al > Mn > Zn > Cr > Fe > Co > Ni > Sn > Pb > H > Cu > Ag > Hg > Pt > Au**  
More active 

→

 less active

Any metal before hydrogen will displace hydrogen gas,  $H_{2(g)}$ , from an acid, and any metal after hydrogen will not displace hydrogen from an acid. For example:

**Active metals** (left to H) :  $Zn_{(s)} + 2HCl_{(aq)} \rightarrow ZnCl_{2(aq)} + H_{2(g)}$   
**Inactive metals** (right to H) :  $Cu_{(s)} + HCl_{(aq)} \rightarrow \text{no reaction (NR)}$

In performing this experiment, you will observe and compare the reaction of some common metals with acids and salt solutions and establish an order of reactivity, from most active to least active.

### PART I: Reactivity of metals with HCl

**Very active metals** will react with strong acids such as hydrochloric acid. An example considers magnesium:

Molecular equation:  $Mg_{(s)} + 2HCl_{(aq)} \rightarrow MgCl_{2(aq)} + H_{2(g)}$

# Activity Series Lab Experiment 15 Answers

**Daniela Niemeyer**



## Activity Series Lab Experiment 15 Answers:

**Scientific and Technical Aerospace Reports** ,1974      **Energy Research Abstracts** ,1990      *Energy Research Abstracts* ,1990      **Physics Laboratory Experiments** Jerry D. Wilson,1994      **How to Write a Lab Report** Jerome N. Borowick,2000 This guide outlines an effective methodology for writing the experimental laboratory report showing how skills that emphasize correct grammar and appropriate style must be adapted to writing reports with a purpose reports that emphasize structure and content to persuade the readers It first covers basic principles then explores each section of a report step by step with sample report sections and critiques The Laboratory Report Writing Process Principles of Clear Lab Report Writing Rules of Practice for Lab Report Writing Graphics The Title Page and Table of Contents The Beginning of the Report The Body of the Report The Ending of the Report A Sample Student Lab Report For anyone who must write lab reports as part of their professional responsibilities      **Electricity, Principles and Applications** Richard J. Fowler,1979

**Math Advantage, Grade 3** Harcourt Brace Publishing,1997-06      Annual Report of the Agricultural Experiment Station, Michigan State University Michigan State University. Agricultural Experiment Station,1927      *Annual Report of the Agricultural Experiment Station of the State Agricultural College of Michigan for the Year Ending June 30* State Agricultural College (Mich.). Agricultural Experiment Station,1927      **U.S. Government Research & Development Reports** ,1967

INIS Atomindex ,1996      *Activities manual for electronics principles and applications* Charles A. Schuler,1979

**Government Reports Announcements & Index** ,1980      *El-Hi Textbooks and Serials in Print* ,1987      Government reports annual index ,199?      **Publications** University of Minnesota. School of Physics and Astronomy,1968      *IGY General Report Series* ,1963      The Environment Index ,1987      Research in Education ,1970      *Bibliography of Scientific and Industrial Reports* ,1967

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