

Answer Key – Class 10 Science: Chemical Reactions and Equations

Section A: Multiple Choice Questions

1. c) $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
2. b) Copper is deposited on iron
3. a) Hydrogen
4. d) All of the above
5. d) $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$

Section B: Very Short Answer Questions

6. $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2 \uparrow$
7. A precipitation reaction is a reaction where an insoluble solid (precipitate) is formed.
Example:
 $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} \downarrow + \text{NaNO}_3$
8. Colour changes from blue to light green. Equation: $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
9. Double displacement reaction.
Balanced: $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 \downarrow + 2\text{NaCl}$
10. (i) Change in colour
(ii) Evolution of gas / Formation of precipitate

Section C: Short Answer Questions

11. A redox reaction involves both oxidation and reduction.
Example: $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
Zn is oxidised (reducing agent), Cu^{2+} is reduced (oxidising agent).
12. On heating, lead nitrate decomposes to give lead oxide, nitrogen dioxide and oxygen.
 $2\text{Pb(NO}_3)_2 \rightarrow 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$
Type: Thermal decomposition reaction.

10 1 Chemical Reactions Answer Key

Thomas Edward Thorpe



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1. Understanding the eBook 10 1 Chemical Reactions Answer Key
 - The Rise of Digital Reading 10 1 Chemical Reactions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 10 1 Chemical Reactions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 10 1 Chemical Reactions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 10 1 Chemical Reactions Answer Key
 - Personalized Recommendations
 - 10 1 Chemical Reactions Answer Key User Reviews and Ratings
 - 10 1 Chemical Reactions Answer Key and Bestseller Lists
5. Accessing 10 1 Chemical Reactions Answer Key Free and Paid eBooks
 - 10 1 Chemical Reactions Answer Key Public Domain eBooks
 - 10 1 Chemical Reactions Answer Key eBook Subscription Services
 - 10 1 Chemical Reactions Answer Key Budget-Friendly Options

6. Navigating 10 1 Chemical Reactions Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 10 1 Chemical Reactions Answer Key Compatibility with Devices
 - 10 1 Chemical Reactions Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 10 1 Chemical Reactions Answer Key
 - Highlighting and Note-Taking 10 1 Chemical Reactions Answer Key
 - Interactive Elements 10 1 Chemical Reactions Answer Key
8. Staying Engaged with 10 1 Chemical Reactions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 10 1 Chemical Reactions Answer Key
9. Balancing eBooks and Physical Books 10 1 Chemical Reactions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 10 1 Chemical Reactions Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 10 1 Chemical Reactions Answer Key
 - Setting Reading Goals 10 1 Chemical Reactions Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 10 1 Chemical Reactions Answer Key
 - Fact-Checking eBook Content of 10 1 Chemical Reactions Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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