

Practice Exam 2

Ch 6 & 7 (Chaps. 1, 2, 3, 4, 5)

1. Which of the following is a physical change?
 (a) rusting of metal (b) fruits ripen (c) baking a cake
 (d) ice melts (e) milk turns sour

2. Which of the following is a homogeneous mixture?
 (a) water and sand (b) sand, alcohol, and sugar
 (c) oil and vinegar (d) milk of magnesium (e) air

3. Which of the following is a correct?
 (a) sodium, S (b) mercury, Mg (c) iron, K, Sn
 (d) potassium, P (e) gold, Au

4. What is the equivalent temperature for 105 Fahrenheit in Kelvin?
 (a) 41 (b) 314 (c) 378 (d) 457 (e) -32

5. A sample of carbon tetrachloride has a mass of 39.75 g and a volume of 25.0 mL. What is its density (g/mL)?
 (a) 1.59 (b) 1.74 (c) 2.36 (d) 2.89 (e) 3.50 $d = \frac{m}{V} = \frac{39.75}{25.0 \text{ mL}}$

6. How many significant figures are there in the measurement of 1.0056×10^7 ?
 (a) 1 (b) 2 (c) 3 (d) 4 (e) 5

7. How many significant figures are appropriate to show in the result after carrying out the operation below?

$$(13.7 + 0.027) + 8.221 \rightarrow \text{X, } \div \text{ (last s.f.) } +, - \text{ (last decimal)}$$

 (a) 1 (b) 2 (c) 3 (d) 4 (e) 5

8. How many protons, electrons, and neutrons are there in $^{120}\text{Ba}^{2+}$?
 (a) 56p, 56e, 130n (b) 56p, 54e, 130n (c) 56p, 58e, 74n
 (d) 56p, 54e, 74n (e) 54p, 56e, 74n

9. Which of the following pairs would have similar chemical and physical properties?
 (a) Na, Mg (b) H, Li (c) He, Ne (d) Cr, Ca (e) Cl, O

10. Which of the following is the molecular formula?
 (a) NaCl (b) CsF (c) HCl (d) AlBr₃ (e) NaCN

11. Which of the following is the empirical formula?
 (a) C₂H₆ (b) CH₃COOH (c) P₂O₅ (d) NH₄NO₃ (e) Ag₂CO₃

12. Which of the following is a correct match?
 (a) BaCl₂, barium dichloride (b) CS₂, carbon sulfide
 (c) CuF₂, copper (II) fluoride (d) HNO₃, nitrous acid
 (e) MgSO₄·7H₂O, magnesium sulfate heptahydrate

13. Which of the following pairs is a correct match?
 (a) Ca₃(PO₄)₂, calcium (III) phosphate
 (b) (NH₄)₂SO₄, ammonium sulfate
 (d) Cu(NO₃)₂, copper dinitrate (c) Mg₂O₃, magnesium oxide
 (e) H₂CO₃, carbon acid

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