

Equilibrium Extra Practice Problems

- Quantities of PCl_3 and Cl_2 were placed in a reaction chamber and heated to 230°C at 101.3 kPa. At equilibrium $[\text{PCl}_3] = 0.235 \text{ M}$ and $[\text{PCl}_5] = 0.174 \text{ M}$. If the equilibrium constant for this reaction is 0.0205, then calculate the $[\text{Cl}_2]$ at equilibrium. (0.0277 M)



- 1.0 mole of NH_3 was introduced into a 1.00 L reaction vessel at a certain temperature. At equilibrium 0.60 mols of H_2 were found to be present. Calculate the equilibrium constant. (0.12)



- Consider the hypothetical reaction: $\text{A(g)} + 2\text{B(g)} \rightleftharpoons 2\text{C(g)}$

The initial concentrations of the reactants and products were as follows: $[\text{A}] = 2.00 \text{ M}$; $[\text{B}] = 3.00 \text{ M}$; $[\text{C}] = 1.00 \text{ M}$. At equilibrium, it was found that the $[\text{C}]$ was 3.00 M. Calculate the equilibrium constant. (9.00)

- Consider the hypothetical reaction: $2\text{A(g)} + \text{B(g)} \rightleftharpoons \text{C(g)} + 3\text{D(g)}$

Given that the initial concentrations of A, B, C & D were 4.0 M, 3.5 M, 2.0 M and 5.5 M respectively, and that $[\text{A}]$ at equilibrium was 6.0 M, calculate the equilibrium constant. (0.096)

- Calculate the equilibrium concentrations of all substances involved in the reaction depicted below when 0.10 M CO(g) is mixed with 0.10 M $\text{H}_2\text{O(g)}$ in a 1L reaction vessel. ($x = 0.067 \text{ M}$)



- The equilibrium constant for the reaction above at 700°C is 5.10. Calculate the equilibrium concentrations of all substances when 1.00 mole of each compound is initially placed in a 1.00 L flask. ($x = 0.396 \text{ M}$)
- The equilibrium constant for the reaction below is 4.0. If 2.00 moles of each substance is placed in a 1.00 L flask, calculate the concentrations of all substances once equilibrium is reached. ($x = 0.500 \text{ M}$)



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Christian Drosten



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