

# Science

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(Chapter – 1) (Chemical Reactions and Equations) (Intext)  
(Class 10)

Page 10

## Question 1:

A solution of a substance 'X' is used for white washing.

(i) Name the substance 'X' and write its formula.

(ii) Write the reaction of the substance 'X' named in (i) above with water.

## Answer 1:

(i) The substance 'X' is calcium oxide. Its chemical formula is CaO.

(ii) Calcium oxide reacts vigorously with water to form calcium hydroxide (slaked lime)



## Question 2:

Why is the amount of gas collected in one of the test tubes in Activity 1.7 double of the amount collected in the other? Name this gas.

## Answer 2:

During the Electrolysis of water, hydrogen and oxygen is get separated by the electricity. Water (H<sub>2</sub>O) contains two parts hydrogen and one part oxygen. Since hydrogen goes to one test tube and oxygen goes to another, the amount of gas collected in one of the test tubes is double of the amount collected in the other.

Page 13

## Question 1:

Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

## Answer 1:

When an iron nail is dipped in a copper sulphate solution, iron (which is more reactive than copper) displaces copper from copper sulphate solution forming iron sulphate, which is green in colour.



Therefore, the blue colour of copper sulphate solution fades and green colour appears.

## Question 2:

Give an example of a double displacement reaction other than the one given in Activity 1.10.

## Answer 2:

Sodium carbonate reacts with calcium chloride to form calcium carbonate and sodium chloride.



In this reaction, sodium carbonate and calcium chloride exchange ions to form two new compounds. Hence, it is a double displacement reaction.

## Question 3:

Identify the substances that are oxidised and the substances that are reduced in the following reactions.

(i)  $4\text{Na(s)} + \text{O}_2\text{(g)} \rightarrow 2\text{Na}_2\text{O(s)}$

(ii)  $\text{CuO(s)} + \text{H}_2\text{(g)} \rightarrow \text{Cu(s)} + \text{H}_2\text{O(l)}$

## Answer 3:

(i) Sodium (Na) is oxidised as it gains oxygen and oxygen gets reduced.

(ii) Copper oxide (CuO) is reduced to copper (Cu) while hydrogen (H<sub>2</sub>) gets oxidised to water (H<sub>2</sub>O).

# 2014 Chemistry Answer Obj And Easy

**Peter Jomo Walla**



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  - The Rise of Digital Reading 2014 Chemistry Answer Obj And Easy
  - Advantages of eBooks Over Traditional Books
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  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
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  - Popular eBook Platforms
  - Features to Look for in an eBook 2014 Chemistry Answer Obj And Easy
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2014 Chemistry Answer Obj And Easy
  - Personalized Recommendations
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  - 2014 Chemistry Answer Obj And Easy and Bestseller Lists
5. Accessing 2014 Chemistry Answer Obj And Easy Free and Paid eBooks
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  - Highlighting and Note-Taking 2014 Chemistry Answer Obj And Easy
  - Interactive Elements 2014 Chemistry Answer Obj And Easy
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  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2014 Chemistry Answer Obj And Easy
- 9. Balancing eBooks and Physical Books 2014 Chemistry Answer Obj And Easy
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2014 Chemistry Answer Obj And Easy
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 2014 Chemistry Answer Obj And Easy
  - Setting Reading Goals 2014 Chemistry Answer Obj And Easy
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 2014 Chemistry Answer Obj And Easy
  - Fact-Checking eBook Content of 2014 Chemistry Answer Obj And Easy
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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