

# Exploring the Periodic Table

## Key Ideas

- Why do elements within a group of the periodic table have similar chemical properties?
- What happens to an atom that gains or loses electrons?
- What are the three main categories of elements?

## Key Terms

- ion
- metal
- nonmetal
- semiconductor

## Why It Matters

The properties of metals make metals useful for conducting electricity. For example, wires that carry electricity are made of metal.

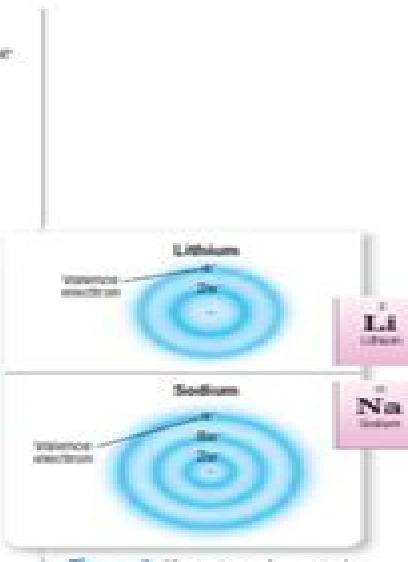
**W**hy is neon an unreactive element? Why is sodium so reactive that it reacts violently with moisture and oxygen in the air? These chemical properties are related to the number of electrons in each element.

### The Role of Electrons

The periodic table is organized by atomic number, which is the number of protons in an atom. For a neutral atom, the number of protons equals the number of electrons. **The periodic trends in the periodic table are the result of electron arrangement.** Specifically, the chemical properties of each group are largely determined by the number of valence electrons. These electrons are closest to the outside of the atom and are sometimes considered part of an outer "shell" of electrons.

#### Valence electrons account for similar properties.

The number of valence electrons determines many of the chemical properties of an element. The diagrams in **Figure 1** show atoms of two elements from Group 1: lithium and sodium. Because they each have one valence electron, lithium and sodium have similar chemical properties. In general, elements in a group have chemical and physical properties in common because they have the same number of valence electrons. Of course, elements in a group are not exactly alike. They differ in numbers of protons in their nuclei and in numbers of electrons in their filled outer energy levels. As a result, the properties of elements in a group are not exactly the same.



**Figure 1** These two elements in Group 1 each have one valence electron. What is the difference in the location of these electrons?

## SECTION 2

### Focus

In this section, students learn about the process of ionization and how valence electrons account for the periodic properties of elements. The section concludes with a brief description of how elements are classified as metals, semiconductors, and nonmetals according to their properties and where, in general, these categories can be found in the periodic table.

### Bellringer

Use the Bellringer transparency to prepare students for this section.

### Demonstrate

**Conductivity** For this demonstration you will need a conductivity tester and samples of elements, such as a copper coin, an iron nail, and pieces of charcoal and sulfur. A conductivity tester can be made from insulated wire, a battery, and a flashlight bulb.

Use the conductivity tester to see which elements conduct an electric current. Ask students: "Which elements are metals?" "How do you know?" (Generally, elements that conduct a current under standard conditions are metals.) "Which elements are nonmetals?" "How do you know?" (Generally, elements that do not conduct a current under standard conditions are nonmetals.)

#### Logical

#### Answer to caption question

Lithium has one valence electron in the second shell, while sodium has one valence electron in the third shell.

### Key Resources

- Teaching Transparencies**  
C12 Ion Formation  
TM16 Elements and Orbitals
- Visual Concept**  
Valence Electrons
- Cross-Disciplinary Worksheets**  
Integrating Biology—The Elements in Your Body

## Answers For Section Exploring The Periodic Table

**Grace Jasmine,Julia Jasmine**

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**Discovering Science Through Inquiry: Matter Kit** Rachel E.

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questions from calculations of chemical equations to organic chemistry Explanations of important terms formulas and

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