

Some of the general equations for the reactions of the Group 7 elements with metals.

7 Complete the word equations for:

- iron + chlorine → iron chloride
- iron + bromine → iron bromide
- sodium + iodine → sodium iodide
- potassium + bromine → potassium bromide
- lithium + chlorine → lithium chloride
- sodium + chlorine → sodium chloride

2 Make up four sentences using the phrases in the table. Each sentence must include a phrase from each column.

There are	weak forces strong forces	between the two atoms of a chlorine molecule. between a chlorine molecule and its neighbours. between the two atoms of an iodine molecule. between an iodine molecule and its neighbours.
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There are weak forces between the I_2 atoms in a iodine molecule. There are strong forces between a iodine molecule and its neighbours between an iodine molecule and its neighbours.

3. For the sentences A–H below:

- Write **1** next to the sentences that are true for Group 1.
- Write **7** next to the sentences that are true for Group 7.
- Write **8** next to the sentences that are true for both Groups 1 and 7.

- A Going down this group, boiling point increases.
- B Going up this group, the reactions of the elements get more vigorous.
- C Going down this group, the number of protons in atoms of the elements increases.
- D Atoms of the elements in this group form molecules made up of two atoms.
- E Going up this group, the colours of the elements 'get lighter'.
- F The elements at the top of this group are gases.
- G The elements in this group conduct electricity.
- H All the atoms of the elements in this group have the same number of electrons in the outermost shell.

7
1
B
4
5
6
7
8
9
B

Write **M** next to the pair of elements that you predict will react most vigorously.
Write **L** next to the pair of elements that you predict will react least vigorously.

- a. sodium and iodine
- b. lithium and iodine
- c. potassium and fluorine
- d. potassium and bromine
- e. sodium and bromine

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Answers For A Ls Chemistry

Marcus Tullius Cicero

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