

1. A gas has a pressure of 0.370 atm at 260.0 °C. What is the pressure at standard temperature?

$$P_2 = P_1 \frac{T_2}{T_1} = \frac{0.370 \text{ atm} (273.15)}{523.15 \text{ K}} = 0.213 \text{ atm}$$

2. If 2.00 moles of gas occupies 4.50L at STP. How much of the same gas will occupy 3.00L at STP?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{2.00 \text{ moles} (22.4 \text{ L})}{4.50 \text{ L}} = 1.00 \text{ mole}$$

3. I've got a car with an internal volume of 12,000 L. If I drive my car into the river and it implodes, what will be the volume of the gas when the pressure goes from 1.0 atm to 1.4 atm?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{1.0 \text{ atm} (12,000 \text{ L})}{1.4 \text{ atm}} = 8640 \text{ L}$$

4. What's the partial pressure of carbon dioxide in a container that holds 5.00 moles of carbon dioxide, 3.00 moles of nitrogen, and 1.00 mole of hydrogen and has a total pressure of 1.05 atm?

$$P_{CO_2} = P_1 \times \frac{\text{moles } CO_2}{\text{Total moles}} = (1.05 \text{ atm}) \times \frac{(5.00 \text{ moles})}{(9.00 \text{ moles})} = 0.555 \text{ atm}$$

5. I have 130 liters of gas in a piston at a temperature of 250. °C. If I cool the gas until the volume decreases to 85 liters, what will temperature of the gas be?

$$T_2 = \frac{P_1 V_1}{P_2 V_2} = \frac{1300 \text{ K} (85 \text{ L})}{130 \text{ L}} = 360 \text{ K} = 70 \text{ °C}$$

6. If a gas in a closed container is pressurized from 150.0 atmospheres to 160.0 atmospheres and its original temperature was 25.0 °C, what would the final temperature of the gas be?

$$T_2 = \frac{P_1 V_1}{P_2 V_2} = \frac{1299.5 \text{ K} (160.0 \text{ atm})}{(150.0 \text{ atm})} = 3.38 \text{ K}$$

7. To compresses nitrogen at 1 atm from 350 mL to 500 mL, what must the new pressure be if the temperature is kept constant?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{1299.5 \text{ K} (350 \text{ mL})}{(500 \text{ mL})} = 2.35 \text{ atm}$$

8. Calcium carbonate decomposes at 1,000. °C to form carbon dioxide and calcium oxide. If 250 liters of carbon dioxide are collected at 1,000. °C, what will the volume of this gas be after it cools to 25. °C?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{1299.5 \text{ K} (250 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

9. 5.00 L of a gas is known to contain 0.265 moles. If this amount of gas is increased to 1.50 moles, what new volume will result (at an unchanged temperature and pressure)?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{1299.5 \text{ K} (1.50 \text{ moles})}{(0.265 \text{ moles})} = 545 \text{ L}$$

## Answer Sheet To Mixed Gas Laws

**Sir Norman Lockyer**

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