

1. Answer: A.

**Explanation:** The value of  $\Delta H$  is negative. Therefore, the reaction is exothermic or heat-releasing.

2. Answer: A.

**Explanation:** The entire reaction was multiplied by  $\frac{1}{2}$ . Therefore,  $\Delta H$  must also be halved.

3. Answer: C.

**Explanation:** Option C is incorrect because  $-566 \text{ kJ/mol}$  is the  $\Delta H$  per mole of  $\text{O}_2$  and not for 2 moles. Even though options B and D are expressed in grams, they are still correct because 56 g of  $\text{CO}$  still corresponds to 2 moles of  $\text{CO}$ , and 32 g of  $\text{O}_2$  still corresponds to 1 mole of  $\text{O}_2$ , which is consistent with the expected factor.

4. Answer: A.

**Explanation:**  $\Delta H = (98 \text{ g } \text{CO}_{(g)}) \left( \frac{1 \text{ mol } \text{CO}_{(g)}}{28 \text{ g } \text{CO}_{(g)}} \right) \left( \frac{-566 \text{ kJ}}{2 \text{ mol } \text{CO}_{(g)}} \right) = -990.5 \text{ kJ}$

5. Answer: B.

**Explanation:**  $\Delta H = (1.5 \text{ mol } \text{CO}_{(g)}) \left( \frac{-566 \text{ kJ}}{2 \text{ mol } \text{CO}_{(g)}} \right) = -424.5 \text{ kJ}$

# Answer Key Unit 6 Thermochemistry

**Charles E. Mortimer**

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