

1. Answer: A

Explanation: The value of ΔH is negative. Therefore, the reaction is exothermic or heat-releasing.

2. Answer: A

Explanation: The entire reaction was multiplied by $\frac{1}{2}$. Therefore, ΔH must also be halved.

3. Answer: C

Explanation: Option C is incorrect because -566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles. Even though options B and D are expressed in grams, they are still correct because 56 g of CO still corresponds to 2 moles of CO, and 32 g of O_2 still corresponds to 1 mole of O_2 , which is consistent with the expected factor.

4. Answer: A

Explanation: $\Delta H = (98 \text{ g CO}_{(g)}) \left(\frac{1 \text{ mol CO}_{(g)}}{28 \text{ g CO}_{(g)}} \right) \left(\frac{-566 \text{ kJ}}{1 \text{ mol CO}_{(g)}} \right) = -990.5 \text{ kJ}$

5. Answer: B

Explanation: $\Delta H = (1.5 \text{ mol CO}_{2(g)}) \left(\frac{-566 \text{ kJ}}{1 \text{ mol CO}_{2(g)}} \right) = -424.5 \text{ kJ}$

Answer Key Unit 6 Thermochemistry

Charles E. Mortimer



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Table of Contents Answer Key Unit 6 Thermochemistry

1. Understanding the eBook Answer Key Unit 6 Thermochemistry
 - The Rise of Digital Reading Answer Key Unit 6 Thermochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Unit 6 Thermochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Unit 6 Thermochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Unit 6 Thermochemistry
 - Personalized Recommendations
 - Answer Key Unit 6 Thermochemistry User Reviews and Ratings
 - Answer Key Unit 6 Thermochemistry and Bestseller Lists

5. Accessing Answer Key Unit 6 Thermochemistry Free and Paid eBooks
 - Answer Key Unit 6 Thermochemistry Public Domain eBooks
 - Answer Key Unit 6 Thermochemistry eBook Subscription Services
 - Answer Key Unit 6 Thermochemistry Budget-Friendly Options
6. Navigating Answer Key Unit 6 Thermochemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Unit 6 Thermochemistry Compatibility with Devices
 - Answer Key Unit 6 Thermochemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Unit 6 Thermochemistry
 - Highlighting and Note-Taking Answer Key Unit 6 Thermochemistry
 - Interactive Elements Answer Key Unit 6 Thermochemistry
8. Staying Engaged with Answer Key Unit 6 Thermochemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Unit 6 Thermochemistry
9. Balancing eBooks and Physical Books Answer Key Unit 6 Thermochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Unit 6 Thermochemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Unit 6 Thermochemistry
 - Setting Reading Goals Answer Key Unit 6 Thermochemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Unit 6 Thermochemistry
 - Fact-Checking eBook Content of Answer Key Unit 6 Thermochemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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