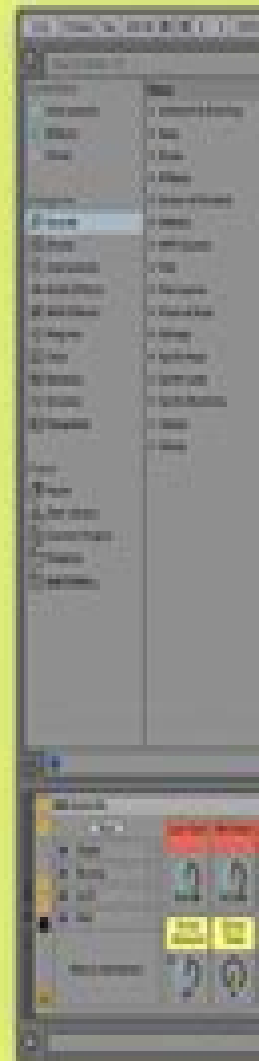
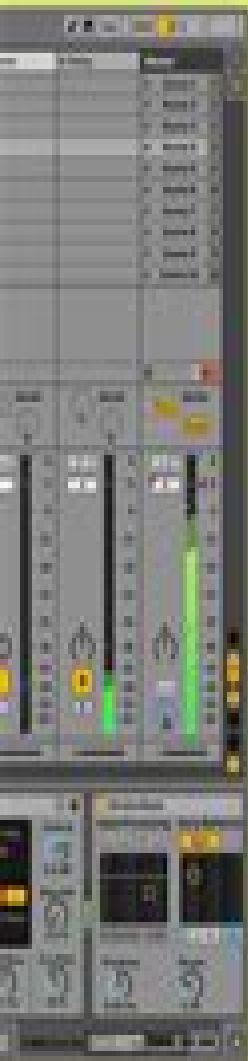


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aqueous solution an example of an arrhenius base is the highly soluble sodium hydroxide NaOH sodium hydroxide dissociates in water as follows $\text{NaOH} \rightarrow \text{Na}^+ + \text{OH}^-$ in water sodium hydroxide

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web acid base reaction theories definitions properties the first attempt at a theoretical interpretation of acid behaviour was made by antoine laurent lavoisier at the end of the 18th century lavoisier supposed that all acids must contain oxygen and this idea was incorporated in the names used for this element in the various languages the english

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web work done by svante arrhenius in the late 19th century set the groundwork for our current understanding of acid base theory despite the relatively limited technology available to him he was able to develop theories that adequately explained the properties and chemical reactivity of many acids and bases

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web the arrhenius theory of acids and bases acids are substances that produce hydrogen ions in solution bases are substances that produce hydroxide ions in solution neutralization happens because hydrogen ions and hydroxide ions react to produce water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$ hydrochloric

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web aug 15 2020 the equilibrium formulated in scheme 1 between hydrogen ion and the corresponding acid and base can be called a simple acid base equilibrium by mixing two simple systems a double acid base system and an acid base equilibrium result that can always be formulated as follows $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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web key points a brønsted lowry acid is any species that is capable of donating a proton H^+ a brønsted lowry base is any species that is capable of accepting a proton which requires a lone pair of electrons to bond to the H^+ water is amphoteric which means it can act as both a brønsted lowry acid and a brønsted lowry base

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