

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---------------------------------|----------------------------------|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ | $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| $pH = -\log[H^+]$ | $pOH = -\log[OH^-]$ |

Part A Completion

Write the appropriate number in the blank space for each item. (Write the number in the blank space.)

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
2. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
3. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
4. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
5. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
6. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
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7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.
9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is _____.
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is _____.

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