

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given pK_a or pK_b
3. Calculate the percent ionization of weak acids

Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant (K_a)
- acid ionization

- base ionization
- base ionization constant (K_b)

Key Equations

$$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \quad \text{Equation 19.1}$$

$$K_b = \frac{[\text{B}][\text{OH}^-]}{[\text{BOH}]} \quad \text{Equation 19.2}$$

Part A Completion

Write the appropriate formula or name for each compound, state whether it is a strong or weak acid or base, and calculate its K_a or K_b value.

1. HNO_3 (strong acid) $K_a = 2.4 \times 10^{-4}$

2. HCl (strong acid) $K_a = 1.3 \times 10^{-2}$

3. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

4. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

5. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

6. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

7. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

8. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

9. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

10. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

11. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

12. HNO_3 (strong acid) $K_a = 2.4 \times 10^{-4}$

13. HCl (strong acid) $K_a = 1.3 \times 10^{-2}$

14. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

15. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

16. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

17. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

18. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

19. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

20. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

21. H_2SO_3 (weak acid) $K_a = 1.0 \times 10^{-2}$

22. H_2SO_4 (strong acid) $K_a = 1.0 \times 10^{-2}$

19 2 Hydrogen Ions And Acidity Section Review

**Kevin T. Patton, Gary A.
Thibodeau, Andrew Hutton**



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