

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~Strontium hydroxide~~ **LiOH**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **LiOH**

HClO₄ is PERMANENTLY acidic & never will completely dissociate in water.

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

HF

HNO₃

Mg(OH)₂

HNO₂

NaOH

NaClO₂

H₂SO₄

3. What is the formula for hydronium? **H₃O⁺**

4. H₂C₂O₄ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for this reaction: **H₂C₂O₄ + H₂O → H₃O⁺ + HC₂O₄⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H₂C₂O₄ is a weak acid.**
- Would you define H₂C₂O₄ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is H₂C₂O₄ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of H₂C₂O₄: **HC₂O₄⁻**
- Identify H₂C₂O₄ as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H⁺ ions.**

5. F⁻ will react with water and gain a H⁺ ion in the process.

- Write a balanced equation for this reaction: **F⁻ + H₂O → HF + OH⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F⁻ is a weak base.**
- Would you define F⁻ as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H⁺ ion.**
- In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F⁻: **HF**
- Identify F⁻ as polyprotic, amphoteric, or neither: **neither**

6. HSO₄⁻ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for this reaction: **HSO₄⁻ + H₂O → H₃O⁺ + SO₄²⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO₄⁻ is a weak acid.**
- Would you define HSO₄⁻ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is HSO₄⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO₄⁻: **SO₄²⁻**
- Identify HSO₄⁻ as polyprotic, amphoteric, or neither: **amphoteric, because it has a negative charge (and will like a proton) and it has H⁺ ions (and will like to lose one).**

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Ibrahim Sikder



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