

**Concept-Development
Practice Page****32-1****Consider this:**

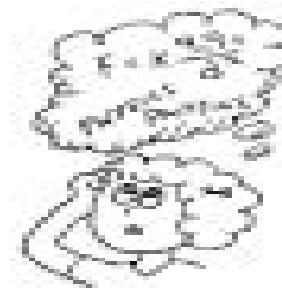
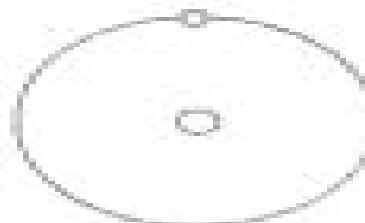
- a. The diagram is of a hydrogen atom.
 - i. Label the proton in the nucleus with a + sign and the orbital electrons with a - sign.
- b. The electrical force between the nucleus and the orbital electrons is a force of
attraction (repulsion).
- c. According to Coulomb's law,

$$F = k \frac{q_1 q_2}{r^2}$$

If the charge of either the nucleus or the orbital electrons were greater, the force between the nucleus and the electron would be
a) greater. b) less.
c) If the distance between the nucleus and electrons were greater than those found in atoms, the
force would be
a) greater. b) less.
c) If the distance between the nucleus and electrons were doubled, the force would be
a) 1/4 as much. b) 1/2 as much. c) 2 times as much. d) 4 times as much.

- d. Consider the electric force between a pair of charged particles a certain distance apart. By Coulomb's law,

- a. If the charge on one of the particles is doubled, the force is
a) quadrupled. b) doubled. c) quadrupled.
d) halved.
- b. If instead the charge on both particles is doubled, the force is
a) quadrupled. b) doubled. c) quadrupled.
d) halved.
- c. If a constant distance between the particles is doubled, the force is
a) quadrupled. b) doubled. c) quadrupled.
d) halved.
- d. If the distance is halved, and the charges of both particles is doubled,
the force is _____ times as great.



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