
AAMC MCAT Practice Exam 1 Questions and Answers

In the chromatography of the reaction mixture, water absorbed on cellulose functioned as the stationary phase. What was the principal factor determining the migration of individual components in the sample?

- A) Hydrogen bonding
- B) Solute concentration
- C) Stationary phase concentration
- D) Thickness of paper - Answer - A) Hydrogen bonding

The answer to this question is A because the relative amount of hydrogen bonding to the stationary phase will determine the relative rate of migration of the various components in the sample.

What assumption is being made if scientists conclude that aspartic acid was formed by the prebiological synthesis in the passage?

- A) Aspartic acid is unstable at temperatures below 150 degrees C.
- B) All of the malic acid underwent the dehydration reaction to form fumaric/maleic acid.
- C) Compound A and cyanide were available on primitive Earth.
- D) The reaction between ammonia and fumaric acid was catalyzed by the presence of water. - Answer - C) Compound A and cyanide were available on primitive Earth.

The answer to this question is C since, in order for the experimental reaction sequence to be relevant to the primordial formation of aspartic acid, the starting materials used (Compound A and cyanide) are assumed to have been available.

According to the developed chromatography plate shown below, what is the approximate R_f value of aspartic acid?

aspartic acid = 2
solvent front = 10

- A) 0.20
- B) 0.50
- C) 5
- D) 10 - Answer - The answer to this question is A (0.20) because R_f is the ratio of the distance travelled by the analyte relative to the solvent front during a chromatographic separation. Aspartic acid travelled two units, while the solvent front travelled ten units, giving an R_f of $2/10 = 0.20$ for aspartic acid.

Which of the following statements does NOT correctly describe the

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Karin Nielsen-Saines



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