
AAMC MCAT Practice Exam 1 Questions and Answers

In the chromatography of the reaction mixture, water absorbed on cellulose functioned as the stationary phase. What was the principal factor determining the migration of individual components in the sample?

- A) Hydrogen bonding
 - B) Solute concentration
 - C) Stationary phase concentration
 - D) Thickness of paper
- Answer - A) Hydrogen bonding

The answer to this question is A because the relative amount of hydrogen bonding to the stationary phase will determine the relative rate of migration of the various components in the sample.

What assumption is being made if scientists conclude that aspartic acid was formed by the prebiological synthesis in the passage?

- A) Aspartic acid is unstable at temperatures below 150 degrees C.
 - B) All of the malic acid underwent the dehydration reaction to form fumaric/maleic acid.
 - C) Compound A and cyanide were available on primitive Earth.
 - D) The reaction between ammonia and fumaric acid was catalyzed by the presence of water.
- Answer - C) Compound A and cyanide were available on primitive Earth.

The answer to this question is C since, in order for the experimental reaction sequence to be relevant to the primordial formation of aspartic acid, the starting materials used (Compound A and cyanide) are assumed to have been available.

According to the developed chromatography plate shown below, what is the approximate R_f value of aspartic acid?

aspartic acid = 2
solvent front = 10

- A) 0.20
 - B) 0.50
 - C) 5
 - D) 10
- Answer - The answer to this question is A (0.20) because R_f is the ratio of the distance travelled by the analyte relative to the solvent front during a chromatographic separation. Aspartic acid travelled two units, while the solvent front travelled ten units, giving an R_f of $2/10 = 0.20$ for aspartic acid.

Which of the following statements does NOT correctly describe the

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 - The Rise of Digital Reading Aamc Practice Test Answer Key Bing
 - Advantages of eBooks Over Traditional Books
2. Identifying Aamc Practice Test Answer Key Bing
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aamc Practice Test Answer Key Bing
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aamc Practice Test Answer Key Bing
 - Personalized Recommendations
 - Aamc Practice Test Answer Key Bing User Reviews and Ratings
 - Aamc Practice Test Answer Key Bing and Bestseller Lists
5. Accessing Aamc Practice Test Answer Key Bing Free and Paid eBooks
 - Aamc Practice Test Answer Key Bing Public Domain eBooks

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- Aamc Practice Test Answer Key Bing Budget-Friendly Options
- 6. Navigating Aamc Practice Test Answer Key Bing eBook Formats
 - ePub, PDF, MOBI, and More
 - Aamc Practice Test Answer Key Bing Compatibility with Devices
 - Aamc Practice Test Answer Key Bing Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aamc Practice Test Answer Key Bing
 - Highlighting and Note-Taking Aamc Practice Test Answer Key Bing
 - Interactive Elements Aamc Practice Test Answer Key Bing
- 8. Staying Engaged with Aamc Practice Test Answer Key Bing
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aamc Practice Test Answer Key Bing
- 9. Balancing eBooks and Physical Books Aamc Practice Test Answer Key Bing
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aamc Practice Test Answer Key Bing
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aamc Practice Test Answer Key Bing
 - Setting Reading Goals Aamc Practice Test Answer Key Bing
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aamc Practice Test Answer Key Bing
 - Fact-Checking eBook Content of Aamc Practice Test Answer Key Bing
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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