

A Voyage through Equations ANSWER KEY

Section 1: Identify the type of reaction

- 1) $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$ DOUBLE DISPLACEMENT
- 2) $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$ DOUBLE DISPLACEMENT
- 3) $\text{C}_6\text{H}_{12} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$ COMBUSTION
- 4) $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ SINGLE DISPLACEMENT
- 5) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ DECOMPOSITION
- 6) $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_5$ SYNTHESIS
- 7) $2 \text{RbNO}_3 + \text{BaF}_2 \rightarrow \text{Ba(NO}_3)_2 + 2 \text{RbF}$ DOUBLE DISPLACEMENT
- 8) $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$ SINGLE DISPLACEMENT
- 9) $\text{C}_2\text{H}_6\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ COMBUSTION
- 10) $2 \text{C}_2\text{H}_6 + \text{Fe} \rightarrow \text{Fe}(\text{C}_2\text{H}_6)_2$ SYNTHESIS
- 11) $\text{SeCl}_4 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$ SINGLE DISPLACEMENT



Section 2: Practicing equation balancing



A Voyage Through Equations Answer Key

Michael Brown

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