



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">hydrogen ionacidbasestrong acidweak acidpH	<ul style="list-style-type: none">hydrogen ion concentrationacidicbasicneutraltitrationtitrant
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Key Equations

- $\text{pH} = -\log[\text{H}^+]$
- $\text{pH} = \log[\text{OH}^-]$

$$\text{pH} = -\log[\text{H}^+]$$

Principles of Acids and Bases

Acids and bases are substances that have the ability to ionize in water to produce hydrogen ions and hydroxide ions, respectively.

Hydrogen ions (H^+) are the smallest hydrogen atoms. They are positively charged because they have lost an electron. They are also called protonic acids because they contain protons.

Hydroxide ions (OH^-) are negatively charged because they have an extra electron. They are also called hydroxide bases because they contain hydroxide ions.

Acids are substances that have the ability to ionize in water to produce hydrogen ions (H^+) and other ions.

Acids are also called protonic acids because they contain protons.

Acids are also called hydroxide bases because they contain hydroxide ions.

Acid-Base Properties

<ul style="list-style-type: none">AcidicBasicNeutralTitrationTitrant	<ul style="list-style-type: none">Acid-Base PropertiesAcid-Base ReactionsAcid-Base IndicatorsAcid-Base NeutralizationAcid-Base Titration
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