

Practice test Acid-Base Equilibrium

1. For the following reaction of hypochlorous acid (HOCl) with water.



what would be the effect of adding sodium hypochlorite (NaOCl) to the reaction at equilibrium?

- a. The concentrations of both HOCl and H_3O^+ would increase.
- b. The concentrations of both HOCl and H_3O^+ would decrease.
- c. The concentration of HOCl would increase and the concentration of H_3O^+ would decrease.
- d. The concentration of HOCl would decrease and the concentration of H_3O^+ would increase.
- e. There would be no change because sodium hypochlorite is a salt without any acidic or basic properties.

2. Calculate the molar hydronium ion concentration in a solution containing 0.23 M hypochlorous acid (HOCl), a monoprotic weak acid used in bleach solutions. For HOCl, $K_a = 2.9 \times 10^{-8}$.

3. Given three separate solutions containing equal concentrations of formic acid ($K_a = 1.7 \times 10^{-4}$), phenol ($K_a = 1.3 \times 10^{-10}$), and acetic acid ($K_a = 1.8 \times 10^{-5}$), select the response below that has the acids arranged in order of increasing percent dissociation at equilibrium.

- a. formic < phenol < acetic
- b. formic < acetic < phenol
- c. acetic < formic < phenol
- d. phenol < acetic < formic
- e. No response is correct.

4. From the following choices, select the one that would be the most basic (least acidic).

- a. 0.1 M hydrochloric acid (a strong acid)
- b. 0.1 M acetic acid (a weak acid)
- c. 0.1 M sodium acetate (the salt of a weak acid)
- d. 0.1 M ammonium chloride (the salt of a weak base)
- e. pure water

Acid Base Equilibrium Practice Test

Frederick J. Rowe

Acid Base Equilibrium Practice Test:

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