

## Acid-Base Titration Lab

### Introduction

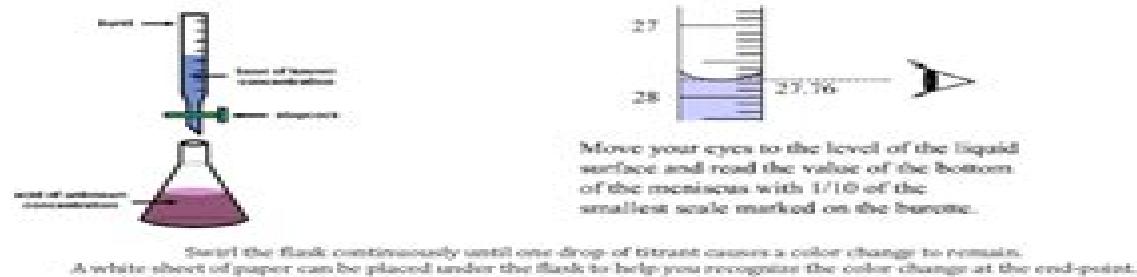
In chemistry laboratory, it is sometimes necessary to experimentally determine the concentration of an unknown acid or base solution. A procedure for making this kind of determination is called an **acid-base titration**. In this laboratory process, a solution of known concentration, called the standard solution, is carefully added to a solution of unknown concentration until the mixture becomes neutral. The neutral point of the solution is recognized by an indicator's color change. If the unknown solution is acidic, then the standard solution will be basic. The opposite would be true if the unknown solution was basic.

We know that the mixing of equal amounts of acid and base ions will create neutral water. At the molecular level, this reaction can be illustrated with the following equation.



This equation states that one mole of hydrogen ions (acid) will neutralize one mole of hydroxide ions (base). Since we can exactly measure the moles of the standard solution, we can assume that the moles of the solution of unknown concentration will be the same at the neutral point. This is called the **end-point** of the titration. Using the equation  $M_1V_1 = M_2V_2$ , we can use the experimental data from the titration to find the unknown concentration.

**Objective:** In this experiment the unknown solution will be HCl(aq) and the standard solution will be the base sodium hydroxide. You will know the concentration of the base and the volume of the acid and base used. With this information you can use the titration formula to calculate the concentration of the acid. The diagram below shows the set up.



### Pre Lab

1. Define the following words:

Titration- \_\_\_\_\_  
Endpoint- \_\_\_\_\_  
Neutralization- \_\_\_\_\_

2. Write the neutralization reaction for HCl reacting with NaOH.
3. What is the pH of the solution at the end point of the titration?

### Procedure

1. The flask is filled with 10 mL of unknown concentration of HCl. Record the volume of acid on your data form
2. Phenolphthalein Indicator is added to the flask.
3. Record on your data form the initial volume of base in the buret.
4. You will now start adding base (0.25M NaOH) from the buret into the flask to neutralize the acid. The flask is gently swirled as the base is added. A pink color should appear as the base is added, but it will disappear as the flask is swirled. The end-point of the titration will be reached when 1 drop of base makes the solution turn and stay pink.
5. Record the final volume of base from the buret that was needed to neutralize the acid.

# Acid Base Titration Lab Answers Ap Chem

**L Towne**

## **Acid Base Titration Lab Answers Ap Chem:**

*CliffsNotes AP Chemistry* Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end. Two full length practice tests with detailed answer explanations are included in the book.

*Illustrated Guide to Home Chemistry Experiments* Robert Bruce Thompson, 2012-02-17 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work. Purify alcohol by distillation. Produce hydrogen and oxygen gas by electrolysis. Smelt metallic copper from copper ore you make yourself. Analyze the makeup of seawater, bone and other common substances. Synthesize oil of wintergreen from aspirin and rayon fiber from paper. Perform forensics tests for fingerprints, blood, drugs and poisons and much more. From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions. But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability. The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab, master laboratory skills and work safely in your lab. The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics: Separating Mixtures, Solubility and Solutions, Colligative Properties of Solutions, Introduction to Chemical Reactions, Stoichiometry, Reduction, Oxidation, Redox Reactions, Acid Base Chemistry, Chemical Kinetics, Chemical Equilibrium and Le Chatelier's Principle, Gas Chemistry, Thermochemistry and Calorimetry, Electrochemistry, Photochemistry, Colloids and Suspensions, Qualitative Analysis, Quantitative Analysis, Synthesis of Useful Compounds, Forensic Chemistry. With plenty of full color illustrations and photos, Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam. A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course. This hands on introduction to real chemistry using real equipment, real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry.

**AP Chemistry** Albert S. Tarendash, Frederick J. Rowe, 2002 This indispensable guide to chemistry helps students who wish to prepare for the AP Chemistry exam on their own. Comprehensive and easy to understand, this learning guide includes a full content review, two full length practice tests with hundreds of practice questions and thorough answer explanations and proven test taking strategies.

**Using Constructivism in Teaching AP Chemistry** Avideh Lotfi, 2004

*Laboratory Experiments for General Chemistry* Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998

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