

## Periodic Trends

**ATOMIC RADIUS**

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
  
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
  
5. Circle the atom in each pair that has the largest atomic radius.  
a) Al B      b) S O      c) Br Cl  
d) Na Al      e) O F      f) Mg Ca

6. Put the following elements in order from smallest to largest atomic radius **and explain why**:  
C, O, Sn, Sr.

**ELECTRONEGATIVITY**

7. Define electronegativity
  
8. How does the ionic radius of a nonmetal compare with its atomic radius?
  
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
  
10. What causes this trend?
  
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
  
12. What causes this trend?
  
  
13. Circle the atom in each pair that has the greater electronegativity.  
a) Ca Ga      b) Li O      c) Cl S      d) Br As      e) Ba Sr      f) O S

**GENERAL QUESTIONS**

14. Which group tends to form +1 ions? \_\_\_\_\_
15. Which group tends to form +2 ions? \_\_\_\_\_
16. Which group tends to form -1 ions? \_\_\_\_\_
17. Which group tends not to form ions or react? \_\_\_\_\_

# Atomic Structure Graphing Periodic Trends Answer Key

**Roman Wölfel**

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