

# APPENDIX G

## Standard Thermodynamic Properties for Selected Substances

Standard Thermodynamic Properties for Selected Substances

Substance	$\Delta H_f^\circ$ (kJ mol <sup>-1</sup> )	$\Delta G_f^\circ$ (kJ mol <sup>-1</sup> )	$S^\circ$ (J K <sup>-1</sup> mol <sup>-1</sup> )
aluminum			
Al(s)	0	0	28.3
Al(g)	324.4	285.7	164.54
Al <sup>3+</sup> (aq)	-531	-485	-321.7
Al <sub>2</sub> O <sub>3</sub> (s)	-1676	-1582	50.92
AlF <sub>3</sub> (s)	-1510.4	-1425	66.5
AlCl <sub>3</sub> (s)	-704.2	-628.8	110.67
AlCl <sub>3</sub> ·6H <sub>2</sub> O(s)	-2691.57	-2269.40	376.56
Al <sub>2</sub> S <sub>3</sub> (s)	-724.0	-492.4	116.9
Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> (s)	-3445.06	-3506.61	239.32
antimony			
Sb(s)	0	0	45.69
Sb(g)	262.34	222.17	180.16
Sb <sub>2</sub> O <sub>3</sub> (s)	-1440.55	-1268.17	220.92
SbCl <sub>3</sub> (g)	-313.8	-301.2	337.80
SbCl <sub>3</sub> (g)	-394.34	-334.29	401.94
Sb <sub>2</sub> S <sub>3</sub> (s)	-174.89	-173.64	182.00
SbCl <sub>5</sub> (s)	-382.17	-323.72	184.10
SbOCl(s)	-374.0	—	—

TABLE G.1

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**Patrick Vollmar**



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